

# Some basic concepts of Chemistry

$$[4.2 \text{ J} = 1 \text{ cal}]$$

- Scientific notation -  $N \times 10^n$
- Accuracy & precision - It is closeness of all experimental results of a particular experiment closeness of experimental result to true value.
- Significant figures;  $0.00\boxed{2341}$ ;  $\boxed{100.001}$ ;  $\boxed{2.00}$ ;  $0.00\boxed{20}$ ;  $\boxed{200}$   
Exact nos has infinite no. of significant figures eg, 3000 pens have 00 SF.
- Rounding off:  $10.5678 \rightarrow 10.568$  @  $10.5674 \rightarrow 10.567$  @  $10.5655 \rightarrow 10.566$  @  $10.5645 \rightarrow 10.564$

For addition & subtraction (only min<sup>m</sup> decimal places)  
 \* multiplication & division (only min<sup>m</sup> least SF)

## (Dalton) Law of multiple proportion

- When 2 elements combine to give more than one compound; the masses of one of the two elements combined with fixed mass of another element always bear a simple whole no. ratio.

## Law of reciprocal proportions

- When an element is combined separately with B & C; the ratio of mass of B & C combined with each other can be equal / simple whole no. multiple of the same ratio calculated from compounds of A with B & A with C.



## Law of multiple proportions

### Oxides of Nitrogen

- $N_2O$  (Nitrous oxide)
- $NO$  (Nitric oxide)
- $NO_2$  (Nitrogen dioxide)
- $NO_2$  (Nitrogen dioxide)
- $N_2O_5$  (Dinitrogen pentoxide)

→ Molecular mass  
 sum of mass of all elements present in the molecule (amu or u)

→ Gram molecular mass  
 mass of 1 mole of molecules (g/mol)

Relative atomic mass - total no. of nucleon present in an atom (unitless)

$\mu$  molecular  $\mu = (\text{unitless}) = \frac{\text{mass of a molecule}}{\frac{1}{12} \times \text{mass of one } ^{12}\text{C atom}}$

→ Atomic mass [u or amu]  
 (mass of an atom)

→ Gram atomic mass [g/mol]  
 (mass of 1 mol of atom)

→ Dulong Petit's law: approx atomic mass of metal  $\times$  specific heat = 6.4

→ Average atomic mass (elements having isotopes)

$\bar{A} = \frac{\sum (\text{at. mass} \times \text{abundance})}{\text{total abundance}}$

→ For ionic compound; whole no. ratio of ions that gives neutral unit is called formula unit of the compound.

→ Mass of 1 formula unit = Formula mass of that compound (amu or u)

Glucose = 180 u  
 Sucrose = 342 u

unitless

## (Lavoisier) Law of conservation of mass

- (not applicable for nuclear rxn)  
 → mass can neither be created nor destroyed

## Law of definite proportions

- (not app for isotopes) (Joseph Proust)  
 → A chemical compound always contains exactly the same proportion of elements by mass.

→ Equal volume of all gases contain equal no. of particles at same T & P conditions

## Avogadro's hypothesis

## Gay-Lussac's law of gaseous volume

- When reactants & products involved in a chem rxn are gas; their mole / stoichiometric ratio becomes their vol. ratio at same temp. & pressure condition (eq<sup>n</sup> must be balanced)

how many time an atom is heavier than  $\frac{1}{12}$  mass of one  $^{12}\text{C}$  atom.

Avogram  
 $1 \text{ amu} = \frac{1}{N_A}$   
 $= 1.67 \times 10^{-24} \text{ g}$

calorie unit

# Notes On The Of Deutreonomy Volume 11

**JS Bruner**



## Notes On The Of Deutereonomy Volume 11:

**The Book of Deuteronomy, Chapters 1-11** Bill T. Arnold, 2022-11-01 The book of Deuteronomy can rightly be called a compendium of the most important ideas of the Old Testament So begins this commentary on the book of Deuteronomy which Bill Arnold treats as the heart of the Torah and the fulcrum of the Old Testament crystallizing the themes of the first four books of the Bible and establishing the theological foundation of the books that follow After a thorough introduction that explores these and other matters Arnold provides an original translation of the first eleven chapters of Deuteronomy along with verse by verse commentary with the translation and commentary of the remaining chapters following in a second volume As with the other entries in the New International Commentary on the Old Testament Arnold remains rooted in the book's historical context while focusing on its meaning and use as Christian Scripture today Ideal for pastors students scholars and interested laypersons this commentary is an authoritative yet accessible companion to the book of Deuteronomy

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Schoberg,2014-02-27 The critical question of the relationship between Jesus and Paul has been well established in New Testament studies for over 150 years and it needs to be addressed with all the tools of historical criticism that scholarship has developed So states Gerry Schoberg at the beginning of this new and important work that does just that By coupling comparative instances in the gospels and the works attributed to Paul Schoberg invites the reader to enquire more profoundly than in past studies as to the nature of the relationship between Jesus and the most dramatic of Christian converts At the heart of this study is not only the question of whether the New Testament truly gives a unified vision of the Christian movement but also how the early followers of Jesus felt able to draw such insightful conclusions about him Answering such questions through the study of Jesus and Paul offer an insight as to how one is to make theological sense of the many details experienced from day to day and thereby live a life in Christ **Thomas Paine and America, 1776-1809**

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