

CHAPTER 1

MATTER AND MEASUREMENT

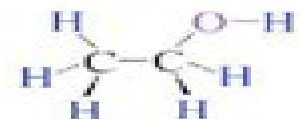
Chemistry: deals with composition, structure, and reactions of matter

- **Matter** – has mass and occupies space
 - Mass – measures quantity of matter
 - Weight – result of gravitational attraction between matter
 - Weight and mass are NOT the same
- **Composition** – what matter is made of and how much of each component is present
 - Expressed by...
 - Weight/mass
 - Volume
 - Percent
 - # of moles
 - # of atoms
 - *Macroscopic* – can be seen and weighed
 - *Submicroscopic level* – number and types of atoms
 - *Atoms* = simple units of matter
 - *Molecules* = combinations of atoms
 - Don't separate when dissolved like compounds → ionic compounds don't exist as molecules and separate when dissolved
 - *Qualitative* = ethanol consists of carbon, hydrogen and oxygen
 - *Quantitative* = ethanol = C₂H₆O
- **Structure** – arrangement of components and how they are bonded
 - Different arrangements of same combination of components = isomers
 - Structures rotate freely
 - Dashed lines = behind dark triangles = forward
- **Reactions** – changes in composition and structure
 - What products?
 - How much of product? (stoich)
 - How fast change occurs?
 - What energy changes accompany rxn?



- exothermic reaction → produces energy over a long period of time

Ethanol
C₂H₆O
or
C₂H₅OH



Scientific Method

- Experiment (record observations) → numerical + observational
 - Careful recordings and analysis of data
 - Controlled conditions so we pinpoint what is causing another thing to happen
 - Reproducible
- Draw a Conclusion – **LAW**
 - Concise statement about a basic relationship or regularity of nature drawn from observations

Notes On General Chemistry

Ellen Z. Hillegass



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1. Understanding the eBook Notes On General Chemistry
 - The Rise of Digital Reading Notes On General Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Notes On General Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Notes On General Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Notes On General Chemistry
 - Personalized Recommendations
 - Notes On General Chemistry User Reviews and Ratings
 - Notes On General Chemistry and Bestseller Lists
5. Accessing Notes On General Chemistry Free and Paid eBooks
 - Notes On General Chemistry Public Domain eBooks
 - Notes On General Chemistry eBook Subscription Services
 - Notes On General Chemistry Budget-Friendly Options
6. Navigating Notes On General Chemistry eBook Formats

- ePub, PDF, MOBI, and More
 - Notes On General Chemistry Compatibility with Devices
 - Notes On General Chemistry Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Notes On General Chemistry
 - Highlighting and Note-Taking Notes On General Chemistry
 - Interactive Elements Notes On General Chemistry
 8. Staying Engaged with Notes On General Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Notes On General Chemistry
 9. Balancing eBooks and Physical Books Notes On General Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Notes On General Chemistry
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Notes On General Chemistry
 - Setting Reading Goals Notes On General Chemistry
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Notes On General Chemistry
 - Fact-Checking eBook Content of Notes On General Chemistry
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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