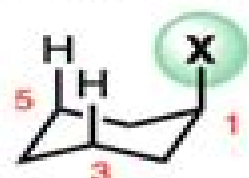


Fill out the values in the following table for various substituents **X**, given either the A value (ΔG), equilibrium constant K or equilibrium ratio of equatorial to axial products at 298 K.



Axial



Equatorial (favored)

Substituent (X)	ΔG (kcal/mol)	K	Equatorial: axial
—CH(CH ₃) ₂	2.15	37.7	97 : 3
—OH	0.87	4.3	81 : 19
—Br	0.43	2.0	67 : 33

$$K = e^{(-\Delta G/RT)}$$

$$\Delta G = -RT \ln K$$

$$T = 298 \text{ K}$$

$$K = \frac{[\text{Products}]}{[\text{Reactants}]}$$

$$(R = 0.0019872 \text{ kcal K}^{-1} \text{ mol}^{-1})$$

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RJ Shavelson



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