

# CHEMISTRY REVISION NOTES

## CONSERVATION OF ENERGY

Energy being a measure of the things that can do work, it is conserved in any process.

The law of conservation of energy states that the total energy of a system is constant.

Energy is a scalar, but it can be split into different forms. Energy is the ability to do work. It is the ability to transfer energy from one form to another. It is the ability to transfer energy from one form to another. It is the ability to transfer energy from one form to another.

Energy can be stored in different ways.

It is the ability to transfer energy from one form to another. It is the ability to transfer energy from one form to another. It is the ability to transfer energy from one form to another.

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## PHASES OF MATTER

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There are four main phases of matter: solid, liquid, gas, and plasma.

Solid: particles are closely packed and vibrate in place.

Liquid: particles are close together but can move past each other.

Gas: particles are far apart and move rapidly in all directions. Plasma: a state of matter consisting of ionized gas.

Changes between phases are called phase transitions. These include melting, freezing, boiling, condensation, sublimation, and deposition.

Phase diagrams show the conditions of temperature and pressure under which different phases of matter exist.

Phase diagrams are used to predict the phase of a substance at a given temperature and pressure. They also show the conditions under which phase transitions occur.

### PHASES

Phase is the state of matter. It is determined by the temperature and pressure of the system.

Phase transitions are changes in the state of matter. They occur when the temperature or pressure of a system changes.

Phase diagrams are used to predict the phase of a substance at a given temperature and pressure.

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## CHEMICAL REACTIONS

A chemical reaction is a process in which one or more substances are converted into one or more different substances.

Chemical reactions are represented by chemical equations.

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