

Thermal Physics

Internal Energy: the sum of the microscopic random kinetic energies and the microscopic potential energies of the molecules of the system.

$$U = \text{sum of microscopic KE} + \text{sum of microscopic PE}$$

Temperature: measure of the average kinetic energies of the molecules that make up the body.

Increase in internal KE = increase in temperature

But increase in $U \neq$ increase in T

Increase in PE = NO increase in temperature!

Thermal Equilibrium: Two objects are said to be in thermal equilibrium if they are in thermal contact and there is no net exchange of thermal energy between them.

In thermal eqm = same temperature

HEAT: the transfer of energy between a system and its environment due to a difference in temperature between them.

ZEROth LAW OF THERMODYNAMICS: If objects A and B are separately in thermal eqm with a third body C, then A and B are also in thermal eqm with each other.

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