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web solution $v_1 = 2.50 \text{ l}$, $v_2 = 2.73 \text{ l}$, $T_1 = 303 \text{ K}$, $T_2 = 273 \text{ K}$ we will be applying Charles law as pressure is constant here $v_1/T_1 = v_2/T_2$ or $v_2 = v_1 \cdot T_2/T_1 = 2.50 \text{ l} \cdot 303 \text{ K} / 343 \text{ K} = 2.18 \text{ l}$ at 90°C a helium sample has a volume of 500 ml determine the temperature at which the volume of the liquid will become 240 ml