

QUANTUM MECHANICS

I. HEISENBERG UNCERTAINTY PRINCIPLE

This principle is a direct consequence of the dual nature of matter, and is also known as 'The Principle of indeterminacy'.

Statement:

"In any simultaneous determination of the position and momentum of the particle, the product of the uncertainty in the x - coordinate of the particle position (Δx) in motion and the uncertainty in the x - component of the momentum (Δp_x) is of the order of or greater than $h/4\pi$ ".

$$\Delta x \Delta p_x \geq \frac{h}{4\pi}$$

→ (1)

Where $\Delta \rightarrow$ minimum uncertainty.

Lower localization

Higher localization

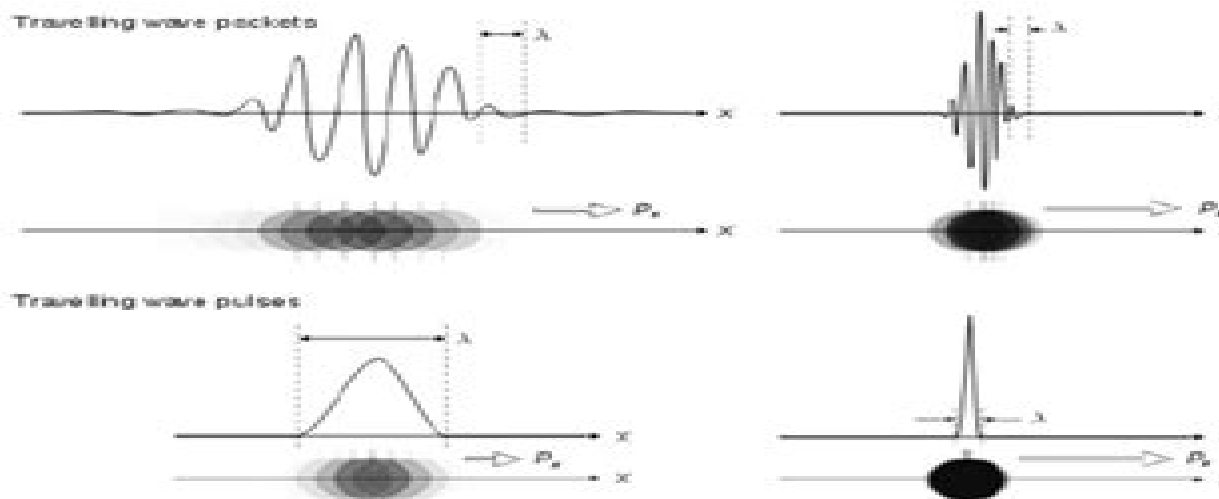


Figure - 1

Explanation:

According to the principle, 'If the momentum of the particle (ex: electron) is known then by the relation $\lambda = h/p$ the wavelength of the associated wave has a unique value i.e. wave is monochromatic. In such cases, the wave packet will have the infinite width and the velocity spread is small so that particle velocity can be determined accurately but the position of the particle may be anywhere between $+\infty$ to $-\infty$ i.e., particle position is uncertain. If the position of the particle is fixed by making the wave packet very small then the wavelength will be too

Quantam Mechanics A First Course

David Kirk



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