

TABLE A-25 Thermochemical Properties of Selected Substances at 298K and 1 atm

Substance	Formula	Molar Mass, M (kg/kmol)	Enthalpy of Formation, h_f° (kJ/kmol)	Gibbs Function of Formation, $\bar{g}_f^\circ$ (kJ/kmol)	Absolute Entropy, $\bar{s}^\circ$ (kJ/kmol · K)	Heating Values	
						Higher, HHV (kJ/kg)	Lower, LHV (kJ/kg)
Carbon	C(s)	12.01	0	0	5.74	32,770	32,770
Hydrogen	H ₂ (g)	2.016	0	0	130.57	141,780	119,950
Nitrogen	N ₂ (g)	28.01	0	0	191.50	—	—
Oxygen	O ₂ (g)	32.00	0	0	205.03	—	—
Carbon monoxide	CO(g)	28.01	−110,530	−137,150	197.54	—	—
Carbon dioxide	CO ₂ (g)	44.01	−393,520	−394,380	213.69	—	—
Water	H ₂ O(g)	18.02	−241,820	−228,590	188.72	—	—
Water	H ₂ O(l)	18.02	−285,830	−237,180	69.95	—	—
Hydrogen peroxide	H ₂ O ₂ (g)	34.02	−136,310	−105,600	232.63	—	—
Ammonia	NH ₃ (g)	17.03	−46,190	−16,590	192.33	—	—
Oxygen	O(g)	16.00	249,170	231,770	160.95	—	—
Hydrogen	H(g)	1.008	218,000	203,290	114.61	—	—
Nitrogen	N(g)	14.01	472,680	455,510	153.19	—	—
Hydroxyl	OH(g)	17.01	39,460	34,280	183.75	—	—
Methane	CH ₄ (g)	16.04	−74,850	−50,790	186.16	55,510	50,020
Acetylene	C ₂ H ₂ (g)	26.04	226,730	209,170	200.85	49,910	48,220
Ethylene	C ₂ H ₄ (g)	28.05	52,280	68,120	219.83	50,300	47,160
Ethane	C ₂ H ₆ (g)	30.07	−84,680	−32,890	229.49	51,870	47,480
Propylene	C ₃ H ₆ (g)	42.08	20,410	62,720	266.94	48,920	45,780
Propane	C ₃ H ₈ (g)	44.09	−103,850	−23,490	269.91	50,350	46,360
Butane	C ₄ H ₁₀ (g)	58.12	−126,150	−15,710	310.03	49,500	45,720
Pentane	C ₅ H ₁₂ (g)	72.15	−146,440	−8,200	348.40	49,010	45,350
Octane	C ₈ H ₁₈ (g)	114.22	−208,450	17,320	463.67	48,260	44,790
Octane	C ₈ H ₁₈ (l)	114.22	−249,910	6,610	360.79	47,900	44,430
Benzene	C ₆ H ₆ (g)	78.11	82,930	129,660	269.20	42,270	40,580
Methanol	CH ₃ OH(g)	32.04	−200,890	−162,140	239.70	23,850	21,110
Methanol	CH ₃ OH(l)	32.04	−238,810	−166,290	126.80	22,670	19,920
Ethanol	C ₂ H ₅ OH(g)	46.07	−235,310	−168,570	282.59	30,590	27,720
Ethanol	C ₂ H ₅ OH(l)	46.07	−277,690	−174,890	160.70	29,670	26,800

Source: Based on JANAF Thermochemical Tables, NSRDS-NBS-37, 1971; Selected Values of Chemical Thermodynamic Properties, NBS Tech. Note 270-3, 1968; and API Research Project 44, Carnegie Press, 1953. Heating values calculated.

Selected Values Of Chemical Thermodynam

Yijin Wang

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Table of Contents Selected Values Of Chemical Thermodynam

1. Understanding the eBook Selected Values Of Chemical Thermodynam
 - The Rise of Digital Reading Selected Values Of Chemical Thermodynam
 - Advantages of eBooks Over Traditional Books
2. Identifying Selected Values Of Chemical Thermodynam
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Selected Values Of Chemical Thermodynam
 - User-Friendly Interface
4. Exploring eBook Recommendations from Selected Values Of Chemical Thermodynam
 - Personalized Recommendations
 - Selected Values Of Chemical Thermodynam User Reviews and Ratings
 - Selected Values Of Chemical Thermodynam and Bestseller Lists
5. Accessing Selected Values Of Chemical Thermodynam Free and Paid eBooks
 - Selected Values Of Chemical Thermodynam Public Domain eBooks
 - Selected Values Of Chemical Thermodynam eBook Subscription Services
 - Selected Values Of Chemical Thermodynam Budget-Friendly Options
6. Navigating Selected Values Of Chemical Thermodynam eBook Formats
 - ePub, PDF, MOBI, and More
 - Selected Values Of Chemical Thermodynam Compatibility with Devices
 - Selected Values Of Chemical Thermodynam Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Selected Values Of Chemical Thermodynam
 - Highlighting and Note-Taking Selected Values Of Chemical Thermodynam
 - Interactive Elements Selected Values Of Chemical Thermodynam
8. Staying Engaged with Selected Values Of Chemical Thermodynam

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Selected Values Of Chemical Thermodynam
- 9. Balancing eBooks and Physical Books Selected Values Of Chemical Thermodynam
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Selected Values Of Chemical Thermodynam
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Selected Values Of Chemical Thermodynam
 - Setting Reading Goals Selected Values Of Chemical Thermodynam
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Selected Values Of Chemical Thermodynam
 - Fact-Checking eBook Content of Selected Values Of Chemical Thermodynam
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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