

For the HF molecule (assuming that it is formed from 1s orbital of H and 2p_z orbital of F). Which of the following is/are correct?
(consider H atom to be represented by A; and F by B)

- ☹ If HF is purely covalent, the valence bond wave function is written as

$$\psi_{\text{covalent}} = \phi_A(1) \phi_B(2) + \phi_A(2) \phi_B(1)$$

- ☹ If HF is purely ionic,

$$\psi_{\text{ionic}} = \phi_A(1) \phi_A(2)$$

- ☹ If HF is purely ionic,

$$\psi_{\text{ionic}} = \phi_B(1) \phi_B(2)$$

- ☹ If HF is 80% covalent and 20% ionic, the wave function is

$$\psi = 0.89 [\phi_A(1) \phi_B(2) + \phi_A(2) \phi_B(1)] + 0.45 [\phi_B(1) \phi_B(2)]$$

3

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Quantum Theory Of Valency

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Quantum Theory Of Valency:

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