

# Examples of Ions

CHEMICAL NAME	FORMULA	BASIC ION	ACIDIC ION	CHEMICAL NAME	FORMULA	BASIC ION	ACIDIC ION
Sodium Chloride	NaCl	$\text{Na}^+$	$\text{Cl}^-$	Lead Bromide	$\text{PbBr}_2$	$\text{Pb}^{2+}$	$\text{Br}^-$
Aluminum Chloride	$\text{NH}_4\text{Cl}$	$\text{NH}_4^+$	$\text{Cl}^-$	Aluminum Sulfide	$\text{Al}_2\text{S}_3$	$\text{Al}^{3+}$	$\text{S}^{2-}$
Copper Oxide	$\text{CuO}$	$\text{Cu}^{2+}$	$\text{O}^{2-}$	Copper Chloride	$\text{CuCl}$	$\text{Cu}^+$	$\text{Cl}^-$
Calcium Oxide	$\text{CaO}$	$\text{Ca}^{2+}$	$\text{O}^{2-}$	Ferric Sulfide	$\text{FeS}$	$\text{Fe}^{2+}$	$\text{S}^{2-}$
Calcium Nitride	$\text{Ca}_3\text{N}_2$	$\text{Ca}^{2+}$	$\text{N}^{3-}$	Calcium Sulfite	$\text{CaSO}_3$	$\text{Ca}^{2+}$	$\text{SO}_3^{2-}$
Magnesium Oxide	$\text{MgO}$	$\text{Mg}^{2+}$	$\text{O}^{2-}$	Magnesium Hydroxide	$\text{Mg}(\text{OH})_2$	$\text{Mg}^{2+}$	$\text{OH}^-$
Zinc Oxide	$\text{ZnO}$	$\text{Zn}^{2+}$	$\text{O}^{2-}$	Lead Sulfate	$\text{PbSO}_4$	$\text{Pb}^{2+}$	$\text{SO}_4^{2-}$
Calcium Carbonate	$\text{CaCO}_3$	$\text{Ca}^{2+}$	$\text{CO}_3^{2-}$	Potassium Nitrate	$\text{KNO}_3$	$\text{K}^+$	$\text{NO}_3^-$

## Simple Ions

$\text{Na}^+$ ,  $\text{Cl}^-$ ,  $\text{Ca}^{2+}$ ,  $\text{Mg}^{2+}$ ,  $\text{Zn}^{2+}$ ,  $\text{Pb}^{2+}$ ,  $\text{Al}^{3+}$ ,  $\text{Cu}^+$ ,  $\text{Fe}^{2+}$

## Compound Ions

$\text{NO}_3^-$ ,  $\text{CO}_3^{2-}$ ,  $\text{NH}_4^+$ ,  $\text{OH}^-$ ,  $\text{SO}_4^{2-}$ ,  $\text{SO}_3^{2-}$

# Radicals Ions And Tibue Damage

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