

Chemistry 200: General Chemistry I

Practice Problems for Final Exam

1. If 4.67 kJ of heat is applied to a mixture of 10.0 g ice ($\Delta H_{fus} = 6.02$ kJ/mole) and 25.0 g water ($C_s = 4.18$ J/g \cdot° C) initially at 0.0 $^\circ$ C, what is the final temperature of the mixture?

2. Oxygen difluoride reacts with water to form hydrofluoric acid and molecular oxygen. Using Table 9.3 on p. 392 in your textbook, calculate ΔH_{rxn} .

3. Carbon monoxide can be synthesized according to the reaction below. If 22.0 g C and 42.0 g O₂ are sealed in a 20.00 L stainless steel vessel at 800 $^\circ$ C, what are the values of P_{O_2} and P_{CO} after the reaction? Use $R = 0.08206$ L \cdot atm/K \cdot mole.



4. A redox titration using the equation shown below is performed on a 50.0 mL sample of the oxidizing agent. If the titration required 22.1 mL of a 0.125 M solution of the reducing agent, what is the concentration of the oxidizing agent in the 50.0 mL sample?



5. Calculate the difference in energy between the $n = 2$ and $n = 4$ levels in the hydrogen atom. Use $R = 1.097 \times 10^7$ m⁻¹.

6. Draw the molecular orbital diagram for Ne₂⁺. Calculate its bond order. Is Ne₂⁺ stable or unstable?

7. When hydrochloric acid is added to solid potassium sulfite, the products formed are liquid water, gaseous sulfur dioxide, and aqueous potassium chloride. If a student wants to form 100.0 mL of sulfur dioxide under STP conditions, but can only use 20.0 mL of hydrochloric acid, what must the concentration of the hydrochloric acid be? Use $R = 0.08206$ L \cdot atm/K \cdot mole.

8. Diatomic hydrogen is created and used in the series of reactions below. If a researcher wants to produce 10.0 kJ of heat, how many grams of zinc are needed?



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Problems and Solutions Supplement to College Chemistry and General Chemistry, Seventh Editions Henry Fuller Holtzclaw, William Harrison Nebergall, William R. Robinson, 1984 **Problems and Solutions for College Chemistry and General Chemistry** John H. Meiser, 1984-01-01 **Problems and Solutions for General Chemistry and College Chemistry, Fifth Editions by Nebergall, Schmidt and Holtzclaw** John H. Meiser, 1976 **Problems and Solutions for General Chemistry and College Chemistry, Sixth Editions by Nebergall, Holtzclaw, and Robinson** Frederick K. Ault, 1980 *General Chemistry Workbook* Daniel C. Tofan, 2010-07-28 This workbook is a comprehensive collection of solved exercises and problems typical to AP introductory and general chemistry courses as well as blank worksheets containing further practice problems and questions It contains a total of 197 learning objectives grouped in 28 lessons and covering the vast majority of the types of problems that a student will encounter in a typical one year chemistry course It also contains a fully solved 50 question practice test which gives students a good idea of what they might expect on an actual final exam covering the entire material **Survival Guide to General Chemistry** Patrick E. McMahon, Rosemary McMahon, Bohdan Khomtchouk, 2019-02-13 This work evolved over thirty combined years of teaching general chemistry to a variety of student demographics The focus is not to recap or review the theoretical concepts well described in the available texts Instead the topics and descriptions in this book make available specific detailed step by step methods and procedures for solving the major types of problems in general chemistry Explanations instructional process sequences solved examples and completely solved practice problems are greatly expanded containing significantly more detail than can usually be devoted to in a comprehensive text Many chapters also provide alternative viewpoints as an aid to understanding Key Features The authors have included every major topic in the first semester of general chemistry and most major topics from the second semester Each is written in a specific and detailed step by step process for problem solving whether mathematical or conceptual Each topic has greatly expanded examples and solved practice problems containing significantly more detail than found in comprehensive texts Includes a chapter designed to eliminate confusion concerning acid base reactions which often persists through working with acid base equilibrium Many chapters provide alternative viewpoints as an aid to understanding This book addresses a very real need for a large number of incoming freshman in STEM fields

Problem-Solving Raymond Chang, 2005 By Brandon J Cruickshank Northern Arizona University and Raymond Chang is a success guide written for use with General Chemistry It aims to help students hone their analytical and problem solving skills by presenting detailed approaches to solving chemical problems Solutions for all of the texts even numbered problems are included *Physical Chemistry for the Biosciences* Raymond Chang, 2005-02-11 This book is ideal for use in a one semester introductory course in physical chemistry for students of life sciences The author's aim is to emphasize the understanding of physical concepts rather than focus on precise mathematical development or on actual experimental details

Subsequently only basic skills of differential and integral calculus are required for understanding the equations The end of chapter problems have both physiochemical and biological applications **Biology/science Materials** Carolina Biological Supply Company,1991 *Problem Solving in General Chemistry* Ronald A. DeLorenzo,1981 Food Processing Technology United States. Office of Education,1967 To assist school administrators and teachers to plan new programs

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