

Key

1. Assign an oxidation number to each of the following indicated elements:

$\text{Pb}(\text{NO}_3)_4$	Cl_2	ClF	LiCl	$\text{Cr}(\text{NO}_3)_3$	CoBr_2	N_2O_5
Pb <u>+4</u>	Cl <u>0</u>	Cl <u>+1</u>	Li <u>+1</u>	Cr <u>+3</u>	Co <u>+2</u>	N <u>+5</u>
N <u>+3</u>		F <u>-1</u>	Cl <u>-1</u>	N <u>+5</u>	Br <u>-1</u>	
MnO_2	Mn_2O_3	KMnO_4	CF_2Cl_2	$\text{Pd}(\text{OH})_2$	CsClO_3	
Mn <u>+4</u>	Mn <u>+3</u>	Mn <u>+7</u>	C <u>+4</u> F <u>-1</u>	Pd <u>+2</u> O <u>-2</u>	Cs <u>+1</u> Cl <u>+5</u>	
O <u>-2</u>		K <u>+1</u>	Cl <u>-1</u>	H <u>+1</u>		

2. Consider the following oxidation-reduction reaction.



Write the two half-reactions. Label each half-reaction as an oxidation or a reduction.



Which species is oxidized?

C in CO

What is the reducing agent?

C in CO

3. In the smelting of iron from iron oxide according to the equation

(a) C in CO (b) O in Fe_2O_3 (c) C in CO_2 (d) Fe metal (e) Fe in Fe_2O_3 4. Identify and justify the Bronsted-Lowry base in the reaction: $\text{NH}_3 + \text{H-Br} \rightarrow \text{NH}_4^+ + \text{Br}^-$ (a) NH_3 because it is a proton donor (b) HBr because it is a proton donor (c) NH_4^+ because it is a proton acceptor (d) HBr because it is a proton acceptor5. A solution is prepared by dissolving 1.28 g of NaNO_3 in enough water to make 500 mL of solution. What is the concentration of the solution in molarity? (The unit of molarity is moles / liter given the symbol M)

$$\begin{aligned} & 22.9898 \\ & + 14.0067 \\ & \times 3 \\ & \hline & 84.9947 \end{aligned}$$

$$\frac{1.28 \text{ g NaNO}_3}{84.9947 \text{ g NaNO}_3} \times \frac{1 \text{ mol NaNO}_3}{1 \text{ mol NaNO}_3} = 0.0151 \text{ mol NaNO}_3$$

$$\frac{0.0151 \text{ mol}}{0.5 \text{ L}} = 0.0302 \text{ M}$$

6. What is the weight (in grams) of NaNO_3 in a 0.500 L bottle of 2.15 M NaNO_3 ?

$$\frac{2.15 \text{ mol NaNO}_3}{1 \text{ L}} \times 0.5 \text{ L} \times \frac{84.9947 \text{ g NaNO}_3}{1 \text{ mol NaNO}_3} = 91.4 \text{ g NaNO}_3$$

Problems For General Chemistry Qualita

M Walker



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