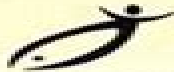


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web the answer is rather simple if you understand electron configurations the shape of the periodic table mimics the filling of the subshells with electrons let us start with h and he their electron configurations are $1s^1$ and $1s^2$ respectively with he the $n=1$ shell is filled

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web elements in the second row of the periodic table place their electrons in the $2n$ shell as well as the $1n$ shell for instance lithium li has three electrons two fill the $1s$ orbital and the third is placed in the $2s$ orbital giving an electron configuration of $1s^2 2s^1$

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web one more point needs to be emphasized about the relationship between electron configuration and the periodic table the atoms of elements in the same vertical column of the table have similar electron configurations for example consider the alkaline earth elements group IIA

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web the electron removed during the ionization of beryllium $He\ 2s^2$ is an s electron whereas the electron removed during the ionization of boron $He\ 2s^2\ 2p^1$ is a p electron this results in a lower first ionization energy for boron even though its nuclear charge is greater by one proton

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web predictions from reliable sources have been used for these elements grayed out electron numbers indicate subshells filled to their maximum bracketed noble gas symbols on the left represent inner configurations that are the same in each period written out these are $He\ 2$ helium $1s^2$ $Ne\ 10$ neon $1s^2\ 2s^2\ 2p^6$

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web q1 what is the electron configuration of the element potassium answer the atomic number of potassium is 19 a neutral potassium element has 19 electrons its electron configuration is $1s^2\ 2s^2\ 2p^6\ 3s^2\ 3p^6\ 4s^1$ or simply put $Ar\ 4s^1$ q2 what is the total number of electrons that can be filled in s p and d subshells

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