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Frank R. Stockton

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web thermochemical equations equations that represent both mass and enthalpy changes $H_2O(s) \rightarrow H_2O(l)$ Δh 6.01 kJ/mol this is an endothermic process it requires 6.01 kJ to melt one mole of ice $H_2O(s)$ the enthalpy value will change if the number of moles varies from the 1:1 reaction stoichiometry this is an exothermic process

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web we will revisit this today and go on to see that if we use a special kind of thermochemical reaction called the standard enthalpy of formation we can calculate enthalpies of reactions without having to manipulate a series of individual thermochemical equations for each step

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web we can express the first law in many ways one of the more useful expressions is that the change in internal energy ΔE of a system in any process is equal to the heat q added to the system plus the work w done on the system by its surroundings $\Delta E = q + w$
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