

Chemistry 102 KEY Practice Exam 1 (Problems 14-15)

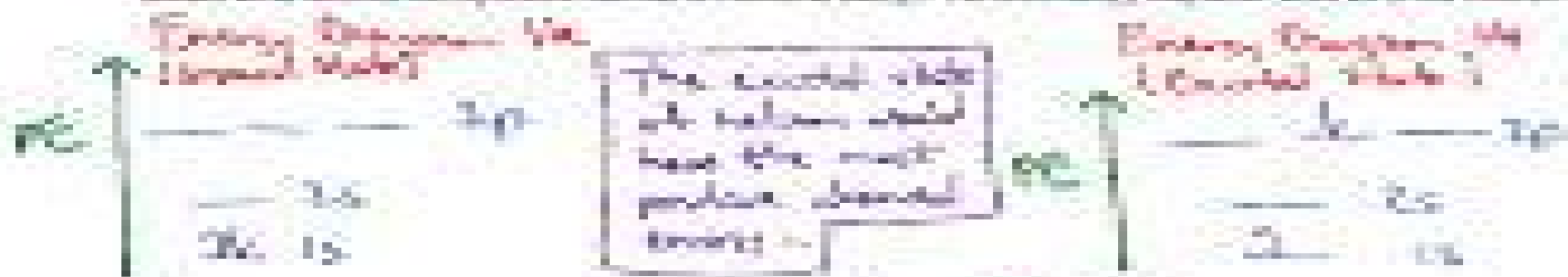
14. Write the **valence** electron configuration for the ground state of an **iron atom**.

Remember: **valence** refers to the

5s, 3d electrons



15. Draw two separate energy diagrams for the following ions, one energy diagram describing the ground state, and the excited energy diagram describing an excited state. Write only the **y-axis** labels and the energy diagram labels. A **boxed** word indicates a subshell containing a certain number of electrons. **Do not** include energy. The **excited-state** diagram, which will provide evidence,



16. Draw the orbital diagram for the ground state **iron(II) ion**, **Fe²⁺**, and the **excited** **Fe²⁺** ion. Do not include subshell labels.



17. The **ionization energy** (the minimum energy required to remove an electron from a neutral atom) for the **iron(II) ion** is **1561 kJ/mol**. What is the **wavelength** (in nm) of the **photon** that has the **minimum** energy to remove an electron from the **iron(II) ion**?

$$\lambda = \frac{c}{\nu} = \frac{(3.00 \times 10^8 \text{ m/s})}{(1.561 \times 10^6 \text{ J/mol}) \times 6.022 \times 10^{23} \text{ mol}^{-1}} = 3.19 \times 10^{-10} \text{ m}$$

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Table of Contents Pkg Acp Chemistry 162

1. Understanding the eBook Pkg Acp Chemistry 162
 - The Rise of Digital Reading Pkg Acp Chemistry 162
 - Advantages of eBooks Over Traditional Books
2. Identifying Pkg Acp Chemistry 162
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pkg Acp Chemistry 162
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pkg Acp Chemistry 162
 - Personalized Recommendations
 - Pkg Acp Chemistry 162 User Reviews and Ratings
 - Pkg Acp Chemistry 162 and Bestseller Lists
5. Accessing Pkg Acp Chemistry 162 Free and Paid eBooks
 - Pkg Acp Chemistry 162 Public Domain eBooks
 - Pkg Acp Chemistry 162 eBook Subscription Services
 - Pkg Acp Chemistry 162 Budget-Friendly Options

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 - ePub, PDF, MOBI, and More
 - Pkg Acp Chemistry 162 Compatibility with Devices
 - Pkg Acp Chemistry 162 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Pkg Acp Chemistry 162
 - Highlighting and Note-Taking Pkg Acp Chemistry 162
 - Interactive Elements Pkg Acp Chemistry 162
8. Staying Engaged with Pkg Acp Chemistry 162
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pkg Acp Chemistry 162
9. Balancing eBooks and Physical Books Pkg Acp Chemistry 162
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pkg Acp Chemistry 162
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Pkg Acp Chemistry 162
 - Setting Reading Goals Pkg Acp Chemistry 162
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Pkg Acp Chemistry 162
 - Fact-Checking eBook Content of Pkg Acp Chemistry 162
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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