#### CHEM 112.3: General Chemistry 1

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Textbook: Chemistry, 2019, 6th Edition, by Gilbert et al.

# Chapter 5: Properties of Gases: The Air We Breathe

#### Properties of a Gas

- Neither definite shape nor volume
- Uniformly fills any container
- Exerts pressure on surroundings
- 4. Changes volume with temperature and pressure
- 5. Miscible mixes together in any proportion
- 6. Much less dense than solids or liquids

#### The atmosphere:

- Layer of gases 50 km thick
- Composition is fairly consistent.
- Properties vary with location.
  - Pressure
  - Density

#### TABLE 6.1 Composition of Dry Air"

Component	N. (by volume)
Nitrogen	79.08
Oxygen	20.95
Argon	0.934
Carbon dioxide	0.0400*
Niegm	0.0018
Helitum	0.00052
Methane	0.00048
Hydrogien	0.00000

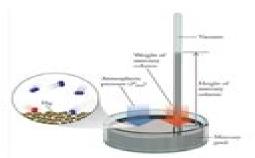
\*Includes major and minor gases (with concentrations > Eppin by volume), "Walve as of May 2015. Americanopheric CO<sub>2</sub> is increasing by about 2 april each year.

#### Pressure (P):

- The ratio of a force (F) to the surface area (A) over which the force is applied
- Atmospheric pressure force exerted by gases surrounding Earth on Earth's surface
- Units of Pressure:
  - SI units = Newton/meter<sup>2</sup> = 1 Pascal (Pa)
  - 1 standard atmosphere (1 atm) = 101,325 Pa
  - 1 atm = 760 mmHg = 760 torr

#### Measurement of Pressure

- Barometer instrument that measures atmospheric pressure
- Height of Hg column based on balance of forces:
  - Gravity (pulls Hg down)
  - Atmospheric pressure (pushes Hg up into evacuated tube)



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# **Table of Contents Pkgacp Cer Chemistry 112**

- 1. Understanding the eBook Pkgacp Cer Chemistry 112
  - The Rise of Digital Reading Pkgacp Cer Chemistry 112
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Pkgacp Cer Chemistry 112
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Pkgacp Cer Chemistry 112
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Pkgacp Cer Chemistry 112
  - Personalized Recommendations
  - Pkgacp Cer Chemistry 112 User Reviews and Ratings
  - Pkgacp Cer Chemistry 112 and Bestseller Lists

- 5. Accessing Pkgacp Cer Chemistry 112 Free and Paid eBooks
  - Pkgacp Cer Chemistry 112 Public Domain eBooks
  - Pkgacp Cer Chemistry 112 eBook Subscription Services
  - Pkgacp Cer Chemistry 112 Budget-Friendly Options
- 6. Navigating Pkgacp Cer Chemistry 112 eBook Formats
  - o ePub, PDF, MOBI, and More
  - Pkgacp Cer Chemistry 112 Compatibility with Devices
  - Pkgacp Cer Chemistry 112 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Pkgacp Cer Chemistry 112
  - Highlighting and Note-Taking Pkgacp Cer Chemistry 112
  - Interactive Elements Pkgacp Cer Chemistry 112
- 8. Staying Engaged with Pkgacp Cer Chemistry 112
  - o Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Pkgacp Cer Chemistry 112
- 9. Balancing eBooks and Physical Books Pkgacp Cer Chemistry 112
  - $\circ$  Benefits of a Digital Library
  - Creating a Diverse Reading Collection Pkgacp Cer Chemistry 112
- 10. Overcoming Reading Challenges
  - $\circ\,$  Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Pkgacp Cer Chemistry 112
  - Setting Reading Goals Pkgacp Cer Chemistry 112
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Pkgacp Cer Chemistry 112
  - Fact-Checking eBook Content of Pkgacp Cer Chemistry 112
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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