

Some basic concepts of Chemistry

$$[4.2 \text{ J} = 1 \text{ cal}]$$

- Scientific notation - $N \times 10^n$
- Accuracy & precision - It is closeness of all experimental results of a particular experiment closeness of experimental result to true value.
- Significant figures; $0.00\boxed{2341}$; $\boxed{100.001}$; $\boxed{2.00}$; $0.00\boxed{20}$; $\boxed{200}$
Exact nos has infinite no. of significant figures eg, 3000 pens have 00 SF.
- Rounding off: $10.5678 \rightarrow 10.568$ @ $10.5674 \rightarrow 10.567$ @ $10.5655 \rightarrow 10.566$ @ $10.5645 \rightarrow 10.564$

For addition & subtraction (only min^m decimal places)
 * multiplication & division (only min^m least SF)

(Dalton) Law of multiple proportion

- When 2 elements combine to give more than one compound; the masses of one of the two elements combined with fixed mass of another element always bear a simple whole no. ratio.

Law of reciprocal proportions

- When an element is combined separately with B & C; the ratio of mass of B & C combined with each other can be equal / simple whole no. multiple of the same ratio calculated from compounds of A with B & A with C.



Law of multiple proportions

Oxides of Nitrogen

- N_2O (Nitrous oxide)
- NO (Nitric oxide)
- NO_2 (Nitrogen dioxide)
- NO_2 (Nitrogen dioxide)
- N_2O_5 (Dinitrogen pentoxide)

→ Molecular mass
 Zat. mass of all elements present in the molecule (amu or u)

→ Gram molecular mass
 mass of 1 mole of molecules (g/mol)

→ Relative atomic mass - total no. of nucleon present in an atom (unitless)

μ molecular $\mu = (\text{unitless}) = \frac{\text{mass of a molecule}}{\frac{1}{12} \times \text{mass of one } ^{12}\text{C atom}}$

→ Atomic mass [u or amu]
 (mass of an atom)

→ Gram atomic mass [g/mol]
 (mass of 1 mol of atom)

→ Dulong Petit's law: approx atomic mass of metal $\times$ specific heat = 6.4

→ Average atomic mass (elements having isotopes)

$L = \frac{\sum (\text{at. mass} \times \text{abundance}) \text{ of each isotope}}{\text{total abundance}}$

→ For ionic compound; whole no. ratio of ions that gives neutral unit is called formula unit of the compound.

→ Mass of 1 formula unit = Formula mass of that compound (amu or u)

Glucose = 180 u
 Sucrose = 342 u

→ unitless

(Lavoisier) Law of conservation of mass

- (not applicable for nuclear rxn)
 → mass can neither be created nor destroyed

Law of definite proportions

- (not app for isotopes) (Joseph Proust)
 → A chemical compound always contains exactly the same proportion of elements by mass.

Avogadro's hypothesis

- Equal volume of all gases contain equal no. of particles at same T & P conditions

Gay-Lussac's law of gaseous volume

- When reactants & products involved in a chem rxn are gas; their mole / stoichiometric ratio becomes their vol. ratio at same temp. & pressure condition (eqⁿ must be balanced)

How many time an atom is heavier than $\frac{1}{12}$ mass of one ^{12}C atom.

Avogram
 $1 \text{ amu} = \frac{1}{N_A}$
 $= 1.67 \times 10^{-24} \text{ g}$

calorie unit

Notes On The Of Deutonomy Volume 11

Horst Dietrich Preuss



Notes On The Of Deutereonomy Volume 11:

Deuteronomy 1-21:9, Volume 6A Duane Christensen, 2018-04-24 The Word Biblical Commentary delivers the best in biblical scholarship from the leading scholars of our day who share a commitment to Scripture as divine revelation This series emphasizes a thorough analysis of textual linguistic structural and theological evidence The result is judicious and balanced insight into the meanings of the text in the framework of biblical theology These widely acclaimed commentaries serve as exceptional resources for the professional theologian and instructor the seminary or university student the working minister and everyone concerned with building theological understanding from a solid base of biblical scholarship Overview of Commentary Organization Introduction covers issues pertaining to the whole book including context date authorship composition interpretive issues purpose and theology Each section of the commentary includes Pericope Bibliography a helpful resource containing the most important works that pertain to each particular pericope Translation the author s own translation of the biblical text reflecting the end result of exegesis and attending to Hebrew and Greek idiomatic usage of words phrases and tenses yet in reasonably good English Notes the author s notes to the translation that address any textual variants grammatical forms syntactical constructions basic meanings of words and problems of translation Form Structure Setting a discussion of redaction genre sources and tradition as they concern the origin of the pericope its canonical form and its relation to the biblical and extra biblical contexts in order to illuminate the structure and character of the pericope Rhetorical or compositional features important to understanding the passage are also introduced here Comment verse by verse interpretation of the text and dialogue with other interpreters engaging with current opinion and scholarly research Explanation brings together all the results of the discussion in previous sections to expose the meaning and intention of the text at several levels 1 within the context of the book itself 2 its meaning in the OT or NT 3 its place in the entire canon 4 theological relevance to broader OT or NT issues General Bibliography occurring at the end of each volume this extensive bibliography contains all sources used anywhere in the commentary

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Deuteronomy 21:10-34:12, Volume 6B Duane

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Old Testament Theology, Volume II Horst Dietrich Preuss, 1996-01-01. In this work, a part of the Old Testament Library series, Horst Preuss provides a comprehensive analysis of the theology of the Old Testament. He focuses on a detailed assessment of Israel's responses to God's acts of election and covenant with them as a people. The Old Testament Library provides fresh and authoritative treatments of important aspects of Old Testament study through commentaries and general surveys. The contributors are scholars of international standing.

The Saturday Review of Politics, Literature, Science and Art, 1912. The Gender Agenda LIS GODDARD, CLARE HENDRY, 2020-05-21. In an age when men can be nurses or stay at home dads and women can be pilots or business managers, there are few areas left where gender alone determines what a person can and cannot do. Yet different models still exist in Christian ministry. Some denominations contend that certain areas of church life should be the preserve of men alone, while others allow full access to all areas for both sexes. But which is right? In this compelling email exchange, Lis Goddard and Clare Hendry search the Scriptures for guidance on the roles of women and men in church leadership today. Against the busy backdrop of everyday life, their conversation covers all the key passages, leaving no tricky verse unexamined. Passionately arguing their respective corners, they pinpoint where they disagree and agree, all the while modelling Christian debate and friendship. Points to ponder helpfully enable readers to explore their own conclusions.

Digest; Review of Reviews Incorporating Literary Digest, 1897. **The Outlook**, 1913. **Catalogue of English Bible Translations** William J. Chamberlin, 1991-12-10. While other Bible catalogs are available, this comprehensive reference book is destined to become the standard in the field. Chamberlin's one-volume work traces the publication history of multiple editions of Bible translations and offers valuable descriptive annotations. The catalog not only includes complete

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