#### Chemistry 2nd Semester Exam Review Sheet: Chapters 9, 10, 11, 12, 13, 14, 15, and 16

#### Chapter 9: Chemical Quantities

- Solve stoichiometry problems using grams →moles → moles → grams calculations. Examples below:
  - Calculate moles/grams of (NH<sub>4</sub>):SO<sub>4</sub> produced when 30.0 moles of NH<sub>3</sub> react with H<sub>2</sub>SO<sub>4</sub>.

2NH3 + H2SO4 → (NH4)2SO4 30.0 mol × moles

30.0 moles x 1mol/2mol = 15.0 moles

15.0 moles (NH4)2SO4 x 132.2 grams/mole = 1983 = 1980 grams (to 3 s.f.)

b. Na<sub>2</sub>CO<sub>3</sub> + 2HNO<sub>2</sub> → 2NaNO<sub>3</sub> + CO<sub>2</sub> + H<sub>2</sub>O. How many moles/grams of sodium nitrate are produced by the complete reaction of 100, grams of sodium carbonate? Moles of nitric acid needed to produce 500.0 grams of sodium nitrate? Be able to calculate grams or moles of any reactant or product given a quantity of reactant or product.

Grams → Moles → Moles → Grams

100. g Na2CO3 x 1 mol/106.0 g x 2 mol NaNO3 / 1 mol Na2CO3 x 85.00 g/mol = 160.4 g

500.0 g NaNO3 x 1 mol/85.00 g x 2 mol / 2 mol x 63.02 g/mol = 370.7 grams

Calculate percent yield of products. Example below using (b) above.

What is the theoretical yield of NaNO<sub>3</sub> if 100.0 grams of Na<sub>2</sub>CO<sub>3</sub> are used in the reaction? What is the percent yield of NaNO<sub>3</sub> if 130.5 grams of NaNO<sub>3</sub> are actually collected?

Theoretical yield = 160.4 grams (answer from above)

Percent Yield = (Actual Yield/Theoretical Yield) x 100 = (130.5 g/160.4 g) x 100 = 81.36%

What is a limiting reactant in a chemical reaction? The reactant that is used up first, thereby limiting the amount of product that can be formed.

#### Chapter 10: Modern Atomic Theory

4. Describe Rutherford's model of the atom. What could it not explain?

Nucleus (very small) consisting of protons with electrons in orbits around the nucleus. Could not explain how the electrons remained in orbit.

- Define electromagnetic radiation, wavelength, and frequency.
- How do atoms gain energy to become excited? Absorb a photon of energy.

.How do they lose energy to return to their ground state? Release a photon of same energy as that absorbed.

 What is the evidence that energy absorbed by the hydrogen atom is quantized? Specific wavelengths of light seen in emission spectra for atoms. (page 285-86)

Describe the Bohr model of the atom. How did it improve on Rutherford's model?

Same as Rutherford with respect to nucleus. Electrons located on orbits at specific distances from nucleus.

9. Draw the orbital diagram (with the arrows) for an element: for example, phosphorus, sodium, and fluorine.

See book, notes, probs 54, 96

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**M** Tight

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#### Second Semester Laboratory Chemistry 11th Ed Book Review: Unveiling the Magic of Language

In an electronic era where connections and knowledge reign supreme, the enchanting power of language has are more apparent than ever. Its capability to stir emotions, provoke thought, and instigate transformation is truly remarkable. This extraordinary book, aptly titled "Second Semester Laboratory Chemistry 11th Ed," written by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound affect our existence. Throughout this critique, we will delve into the book is central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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## **Table of Contents Second Semester Laboratory Chemistry 11th Ed**

- 1. Understanding the eBook Second Semester Laboratory Chemistry 11th Ed
  - The Rise of Digital Reading Second Semester Laboratory Chemistry 11th Ed
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Second Semester Laboratory Chemistry 11th Ed
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - o Features to Look for in an Second Semester Laboratory Chemistry 11th Ed
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Second Semester Laboratory Chemistry 11th Ed
  - Personalized Recommendations
  - Second Semester Laboratory Chemistry 11th Ed User Reviews and Ratings
  - Second Semester Laboratory Chemistry 11th Ed and Bestseller Lists

- 5. Accessing Second Semester Laboratory Chemistry 11th Ed Free and Paid eBooks
  - Second Semester Laboratory Chemistry 11th Ed Public Domain eBooks
  - Second Semester Laboratory Chemistry 11th Ed eBook Subscription Services
  - Second Semester Laboratory Chemistry 11th Ed Budget-Friendly Options
- 6. Navigating Second Semester Laboratory Chemistry 11th Ed eBook Formats
  - ∘ ePub, PDF, MOBI, and More
  - Second Semester Laboratory Chemistry 11th Ed Compatibility with Devices
  - Second Semester Laboratory Chemistry 11th Ed Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - o Adjustable Fonts and Text Sizes of Second Semester Laboratory Chemistry 11th Ed
  - Highlighting and Note-Taking Second Semester Laboratory Chemistry 11th Ed
  - Interactive Elements Second Semester Laboratory Chemistry 11th Ed
- 8. Staying Engaged with Second Semester Laboratory Chemistry 11th Ed
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Second Semester Laboratory Chemistry 11th Ed
- 9. Balancing eBooks and Physical Books Second Semester Laboratory Chemistry 11th Ed
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Second Semester Laboratory Chemistry 11th Ed
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Second Semester Laboratory Chemistry 11th Ed
  - Setting Reading Goals Second Semester Laboratory Chemistry 11th Ed
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Second Semester Laboratory Chemistry 11th Ed
  - Fact-Checking eBook Content of Second Semester Laboratory Chemistry 11th Ed
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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