

Unit 1 Nuclear Chemistry

Size, structure and stability of the nucleus - n/p ratio, packing fraction, mass defect and binding energy - nuclear fission and fusion, atom bombs -hydrogen bomb – radioactivity, alpha, beta particles and gamma radiation - Soddy-Fajan displacement law, half and average life period - Geiger-Muller Counter and Wilson Cloud Chamber, applications of radioactivity - in medicine, agriculture, carbon and fossil dating - isotopes, isobars, isotones, isodiapheres and nuclear isomers - natural and artificial radioactivity, artificial transmutation of elements, induced radioactivity, preparation of transuranic elements, Q values, nuclear coulombic barrier .

Unit 2 Solid State

Crystalline and amorphous solids, isotropy and anisotropy, elements of symmetry in crystal systems indices - Miller indices, space lattice and unit cell, Bravais lattices, the seven crystal systems and their Bravais lattices, X-ray diffraction - Bragg's equation and experimental methods (powder method and rotating crystal technique), types of crystals - molecular, covalent, metallic and ionic crystals - close packing of spheres – hexagonal, cubic and body centered cubic packing, interstices in packing - types of crystals – molecular, covalent, metallic crystals - defects in crystals – stoichiometric, non-stoichiometric, extrinsic and intrinsic defects.

Unit 3 Liquid state

Properties of liquids-viscosity, surface tension, capillary action, evaporation, vapour pressure, boiling point and distillation, heat transfer involving liquids

Unit 4 Acids, Bases and Non-aqueous solvents

Concepts of acids and bases – hard and soft acids and bases - Pearson's concept, HSAB principle and its application - basis for hard - hard and soft - soft interactions - non-aqueous solvents - general characteristics of non-aqueous solvent - melting point, boiling point, latent heat of fusion and vaporization, and dielectric constant - reactions such as complex formation, redox, precipitation and acid base type in non-aqueous solvents like liquid ammonia, liquid SO_2 and liquid HF .

Unit 5 Water Technology

Soft and hard water – Hardness – units of hardness – alkalinity - dissolved oxygen – water for various types of industries – treatment of water by ion exchange process - boiler feed water – boiler compounds – internal and external conditioning - water for drinking - municipal water treatment – desalination by RO and electro dialysis.

TEXTBOOKS:

1. Marion Clyde Day Jr, Joel Selbin, [Harry H Sisler](#), '*Theoretical Inorganic Chemistry*', LLC, 2012
2. F. A. Cotton and G. Wilkinson, '*Advanced Inorganic Chemistry*', 5th edition, John Wiley and Sons, New York, 1987
3. B. R. Puri, L. R. Sharma, Kalia, '*Principles of Inorganic Chemistry*', Vishal Publishing Co., 2008

REFERENCES:

1. H. S. Arniker, '*Essentials of Nuclear Chemistry*', 4th edition, New Age International Publishers, 2005.

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philip khuri hitti was born in the mount lebanon mutasarrafate on 22 june 1886 into a maronite christian family in the village of shemlan some 25 km southeast from beirut up in mount lebanon he was educated at an american presbyterian mission school at suq al gharb and then at the syrian protestant college after graduating in 1908 he taught there before moving to columbia university where he earned his phd in 1915 and taught semitic languages after world war i he

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