

**Experiment 2: Solubility
Report Sheet**Name and formula of salt used: Potassium Chlorate $KClO_3$ Mass of salt used: 4.53 gVolume of solution at 40.0°C: 44 mL

$V_{\text{total water}}$ (mL)	Mass of salt used (g)	Solubility (g salt/kg water)	Temperature of crystallization (°C) Trial 1	Temperature of crystallization (°C) Trial 2
10.00	4.53 g		88	
20.00	4.53 g		53.2	
30.00	4.53 g		39.1	
40.00	4.53 g		29.8	
50.00	4.53 g		22.9	

In the space below show a sample calculation for the solubility. Use dimensional analysis.

Construct a graph showing the dependence of the solubility of this salt on temperature. **DO NOT AVERAGE THE TEMPERATURES. PLOT ALL DATA POINTS.** Draw the best line (ONE line) among the data points. Construct a second curve on the same graph of the standard values of the solubility of the salt over the experimental temperature range.

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