

A zero order homogeneous gas phase reaction with stoichiometry, $A \rightarrow rR$, proceeds in a constant volume bomb, $\pi = 1$ when $t = 0$ and $\pi = 1.5$ when $t = 1$. If the same reaction, same feed composition and initial pressure proceeds in a constant pressure setup, find V at $t = 1$ if $V = 1$ at $t = 0$. (π = total pressure; t = time, V = volume.)

6

Reaction Kinetics Volume 1 Homogeneous Gas

J. Alistair Kerr



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Handbook of Bimolecular and Termolecular Gas Reactions J. Alistair Kerr, 1987-06-30 This two volume set presents gas phase kinetic data published in the lieterature between 1978 and 1982 inclusive The data are organized according to the class of bimolecular or termolecular reactions For each reaction the table entry includes Arrhenius parameters and rate constants experimetnal temperature type of kinetic system and a reference to a set of footnotes containing additional experimental details and any reference reacdion and their rate constants

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Kinetics and Mechanism John W. Moore, Ralph G. Pearson, 1981-09-30 The third edition of a classic text originally by Frost and Pearson that describes the fundamental principles and established practices that apply to the study and the rates and mechanisms of homogeneous chemical reactions in the gas phase and in solution Incorporates new advances made during the past 20 years in the study of individual molecular collisions by molecular beam laser applications to experimental kinetics theoretical treatments of reaction rates and our understanding of the principles that govern rates of reaction in solution Presents numerous examples of the deduction of mechanism from experiment including intimate details such as stereochemistry and the dependence of

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