

Avogadro's Law:

under the same temperature and pressure, same volume of gases will consist of the same number of moles of gases.

Volume ratio = Mole ratio
of gases of gases

E.g. if 3 moles of gas A occupies 12 dm³, then 3 moles of gas B at the same temperature and pressure will also occupy 12 dm³.

number
of particles

$$\begin{array}{c} + \\ 6.02 \times 10^{23} \end{array} \quad \begin{array}{c} \times \\ 6.02 \times 10^{23} \end{array}$$

% PURITY =

$$\frac{\text{mass of PURE}}{\text{mass of IMPURE}} \times 100\%$$

% YIELD =

$$\frac{\text{EXPERIMENTAL mass}}{\text{CALCULATED mass}} \times 100\%$$

Limiting reagent: check when no. of moles of all reactants are given

Qn: 0.200 mol of H₂ was reacted with 0.400 mol of O₂ to produce water. Determine the limiting reagent of the reaction.

Ans: 0.200 mol of H₂ required 0.100 mol of O₂ for complete reaction. However 0.400 mol of O₂ was supplied. Therefore O₂ is in excess and H₂ is the limiting reagent.

volume (dm³)
of gas at r.t.p.

$$\begin{array}{c} + 24 \text{ dm}^3 \\ \times 24 \text{ dm}^3 \end{array}$$

number of
MOLES

$$\begin{array}{c} \times M_r \\ + M_r \end{array}$$

mass (g)
of substance

Dilution:

number of moles stays the same,
total volume of solution increases,
concentration of solution decreases

$$\text{conc. (diluted)} = \text{conc. (original)} = \frac{\text{vol. (original)}}{\text{vol. (new)}}$$

$$\begin{array}{c} + \text{vol (dm}^3\text{)} \\ \times \text{vol (dm}^3\text{)} \end{array} \text{ of solution}$$

conc. of solution
(mol dm⁻³)

$$\begin{array}{c} \times M_r \\ + M_r \end{array}$$

conc. of solution
(g dm⁻³)

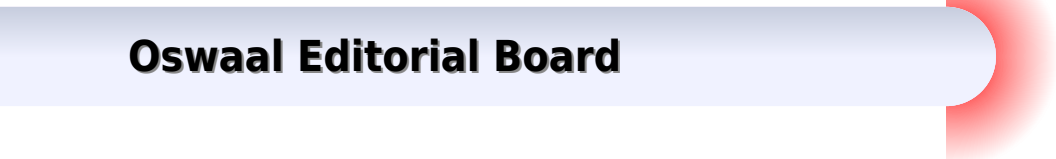
$$\begin{array}{c} + \\ \times \end{array} \text{vol (dm}^3\text{)} \text{ of solution}$$

Sampling:

number of moles decreases,
volume of solution decreases,
concentration remains the same

Mole Concept In Chemistry

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Mole Concept In Chemistry:

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educational failure This book asks the hard questions Are our systems of secondary teaching and the expectations they place on students anachronistic in an age of mass education How far is the curriculum itself responsible for the manifest disparities in achievement between sectors and regions and between boys and girls What has been the universities role in the process of reform and counter reform And what have all these upheavals implied for the practice of teaching **Oswaal JEE**

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