

CHEM 132 - General Chemistry 2

Chapter 11 Notes – "Chemical Kinetics"

Reactions

- Homogenous – reactants and products are in one phase
- Heterogeneous – the reaction takes multiple phases

Speed of Reactions

- Impacted by properties of reactants and products (structure, bonding, etc)
- Concentration of reactants and products
- Temperature – as temp increases, reaction rate increases
- Catalysts – increase reaction rate when added

Rate = (change in reactant/product) divided by (change in time)

- If there is a 1:1 ratio in the reaction, the rate of appearance is equal to the rate of disappearance

Generally, reactions take the following form



Average Rate depends on the interval chosen (represent different slopes)

- Instantaneous Rate – slope of a line tangent to the curve (derivative of a point)

Rate = k (concentration of reactant)

- Rate is proportional to concentration of reactant
- k is rate constant and slope

A general reaction's rate law:

$$\text{Rate} = k[A]^m[B]^n$$

- Where k is rate constant, m and n are order for A and B , and $m+n+\dots$ will represent the overall order of the reaction
 - o Orders are usually integers but they can be fractions

Rate laws MUST be measured and cannot be coming from stoichiometry

- To find the order for a reactant, you must run the experiment with known initial concentrations and measure the initial rate of reaction and the slope at $t=0$

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