

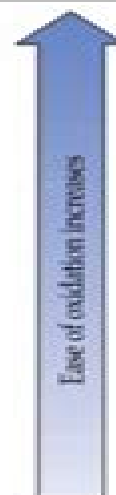


# Chemistry 1411

## EXAM # 2A Sample

Activity Series of Metals in Aqueous Solution

| Metal     | Oxidation Reaction                                       |
|-----------|----------------------------------------------------------|
| Lithium   | $\text{Li}(s) \longrightarrow \text{Li}^+(aq) + e^-$     |
| Potassium | $\text{K}(s) \longrightarrow \text{K}^+(aq) + e^-$       |
| Barium    | $\text{Ba}(s) \longrightarrow \text{Ba}^{2+}(aq) + 2e^-$ |
| Calcium   | $\text{Ca}(s) \longrightarrow \text{Ca}^{2+}(aq) + 2e^-$ |
| Sodium    | $\text{Na}(s) \longrightarrow \text{Na}^+(aq) + e^-$     |
| Magnesium | $\text{Mg}(s) \longrightarrow \text{Mg}^{2+}(aq) + 2e^-$ |
| Aluminum  | $\text{Al}(s) \longrightarrow \text{Al}^{3+}(aq) + 3e^-$ |
| Manganese | $\text{Mn}(s) \longrightarrow \text{Mn}^{2+}(aq) + 2e^-$ |
| Zinc      | $\text{Zn}(s) \longrightarrow \text{Zn}^{2+}(aq) + 2e^-$ |
| Chromium  | $\text{Cr}(s) \longrightarrow \text{Cr}^{3+}(aq) + 3e^-$ |
| Iron      | $\text{Fe}(s) \longrightarrow \text{Fe}^{2+}(aq) + 2e^-$ |
| Cobalt    | $\text{Co}(s) \longrightarrow \text{Co}^{2+}(aq) + 2e^-$ |
| Nickel    | $\text{Ni}(s) \longrightarrow \text{Ni}^{2+}(aq) + 2e^-$ |
| Tin       | $\text{Sn}(s) \longrightarrow \text{Sn}^{2+}(aq) + 2e^-$ |
| Lead      | $\text{Pb}(s) \longrightarrow \text{Pb}^{2+}(aq) + 2e^-$ |
| Hydrogen  | $\text{H}_2(g) \longrightarrow 2\text{H}^+(aq) + 2e^-$   |
| Copper    | $\text{Cu}(s) \longrightarrow \text{Cu}^{2+}(aq) + 2e^-$ |
| Silver    | $\text{Ag}(s) \longrightarrow \text{Ag}^+(aq) + e^-$     |
| Mercury   | $\text{Hg}(l) \longrightarrow \text{Hg}^{2+}(aq) + 2e^-$ |
| Platinum  | $\text{Pt}(s) \longrightarrow \text{Pt}^{2+}(aq) + 2e^-$ |
| Gold      | $\text{Au}(s) \longrightarrow \text{Au}^{3+}(aq) + 3e^-$ |



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