

A2

FOR ADVANCED
D&D™ GAME

Dungeon Module A2 Secret of the Slavers Stockade

by Harold Johnson with Tom Moldvay
AN ADVENTURE FOR CHARACTER LEVELS 4-7



The battle against the slavers continues! You and your fellow adventurers have defeated the slavers of Highport, but you have learned of the existence of another slaver stronghold, and you have decided to continue the attack. But beware! Only the most fearless of adventurers could challenge the slavers on their own ground, and live to tell of it!

This module was originally used for the official ADVANCED DUNGEONS & DRAGONS™ Tournament at Gen Con® XIII Convention and is the second of four in a series of related tournament modules from TSR Hobbies, Inc.

This module contains a challenging scenario, the tournament scoring system plus nine pre-rolled, playtested tournament characters. Also included are large scale referee's maps, notes, and background information. A2 is a complete adventure in itself, but it is also a companion to A1 (Slave Pits of the Undercity), A3 (Assault on the Asile of the Slave Lords) and A4 (In the Dungeons of the Slave Lords). Be sure to look for other exciting adventure modules from the Game Wizards of TSR.

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Secrets Of The Slavers Stockade

Sabine Zange



Secrets Of The Slavers Stockade:

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the boiling point of an organic liquid is the temperature at which its vapour pressure equals the atmospheric pressure over the liquid or it is the temperature at which the vapour and liquid phases are in equilibrium at a given pressure the boiling point is considered as a criterion of purity of a compound and is useful for identification of

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given that the boiling points of ethanol and methanol 78 c and 64 c respectively are far off in a significant way this seems like a valid conclusion conclusion thermometer calibrations are necessary in order to collect reliable information as precise calculations are imperative to gathering correct results

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melting point 1 a push the open end of a capillary tube into the powdered acetamide b move the powder to the closed end of the capillary tube by tapping it on the table repeat until the powdered acetamide occupies 1 2 mm of the capillary tube end c

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discussion and conclusion the technique recrystallization was used because it is a simple method to purify a solid compound the percentage yield for benzoic acid and acetanilide acid being high indicated the initial sample had a low concentration of impurities with the melting point being in a specific

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introduction in this lab we will be using boiling points melting points and infrared spectrometry to identify unknown compounds boiling points are affect by three key factors the first being intermolecular forces inter molecular affect boiling point based on bond strength

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determination of boiling points introduction the boiling point of a compound is the temperature at which it changes from a liquid to a gas this is a physical property often used to identify substances or to check the purity of the compound it

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clean up solid and liquid spills immediately true g conclusions after conducting the boiling point determination lab it can be concluded that the unknown substance 193 is isopropyl alcohol on account of its observed boiling point of 88 c

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conclusion after completing this experiment the melting point of three different unknown compounds was obtained using the mel temp apparatus compound a had a melting point range of 114 6 115 5 c thus revealing it to be acetanilide compound b had a melting point range of 122 6 124 5 c thus

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melting point can be used as the identity and purity of an organic solid however it is not the same as the boiling point the boiling point is a characteristic unique of pure liquids the boiling point is the temperature at which the vapor pressure of a liquid is equal to the atmospheric pressure

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aug 11 2020 boiling points are very sensitive to changes in applied pressure so all boiling points should be reported with the measured pressure a compound s normal boiling point refers to its boiling point at a pressure of 760 text mm ce hg

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the purpose of the experiment is to understand what vapour pressure is and how it changes with temperature to understand

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exploring boiling points abstract in this experiment we are given a set of straight chain alkanes and a set of compounds to compare by following the trendline of the straight chain alkanes it could be concluded that as the molar mass of the straight chain alkanes increases their boiling points increase as well

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laboratory partners samantha and mackenzie goals □ correctly determine the boiling point of an unknown liquid □ measure the refractive index of an unknown □ use the gas chromatogram to determine the mass percentage of a mixture and

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50 °C so in order to convert an experimental measurement taken in Calgary which is at higher altitude than sea level and so is at a lower pressure to that reported for sea level higher pressure one needs to add a correction factor since the boiling point at sea level is higher than that at higher altitudes

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introduction in chemistry one may encounter an unknown substance that needs to be determined by physical properties such as boiling point melting point density solubility etc in this experiment I am to figure out the boiling point of a mystery substance in order to identify it chemical responsibility this experiment requires the handling of a Bunsen burner boiling

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there are a variety of methods by which a sample's boiling point can be determined including distillation reflux and by using a Thiele tube the most straightforward method uses a Thiele tube and has the advantage of using less than 0.5 mL of material

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web may 7 2023 pat the soil down over the newly planted bulb water the bulbs thoroughly cover the area with a few inches of mulch you may omit mulching if you plant the bulbs 12 15 inches deep monitor moisture levels although tiger lilies are fairly drought tolerant they thrive when watered deeply once a week if the blooms look wilted or dry give

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