

WS grade 12 (17.3)

- 1- 1M of nitrogen gas and 1M of hydrogen gas placed in a closed container of volume 2L. Calculate the concentrations of all substances at equilibrium knowing that the concentration of hydrogen gas was found 0.2M at equilibrium, then deduce the value of equilibrium constant.

$$[N_2]_0 = 1M$$

$$[H_2]_0 = 1M$$

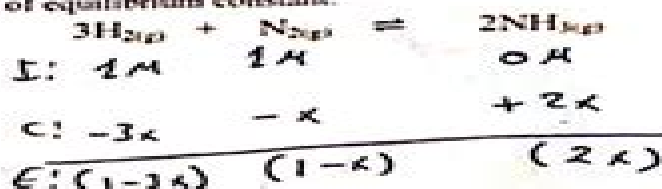
$$V = 2L$$

$$[H_2]_{eq} = 0.2M$$

$$[N_2]_{eq} = ?$$

$$[NH_3]_{eq} = ?$$

$$K_{eq} = ?$$



$$[H_2]_{eq} = 0.2 = 1 - 3x \Rightarrow x = 0.26$$

$$[N_2]_{eq} = 1 - x = 1 - 0.26 = \boxed{0.74M}$$

$$[NH_3]_{eq} = 2x = 2(0.26) = \boxed{0.52M}$$

$$K_{eq} = \frac{[NH_3]^2}{[H_2]^3[N_2]} = \frac{(0.52)^2}{(0.2)^3(0.74)} = \boxed{45.67}$$

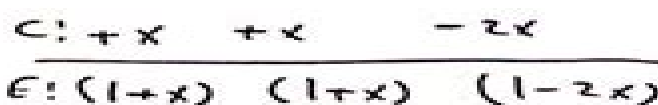
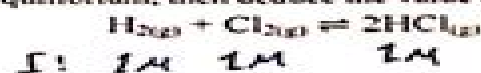
- 2- 1M of HCl, Cl₂, and H₂ placed in a closed container. Calculate the concentrations of all substances at equilibrium knowing that the concentration of hydrogen gas was found 1.3M at equilibrium, then deduce the value of equilibrium constant.

$$[HCl]_0 = 1M$$

$$[Cl_2]_0 = 1M$$

$$[H_2]_0 = 1M$$

$$[H_2]_{eq} = 1.3M$$



$$[H_2]_{eq} = 1.3 = 1 + x \Rightarrow x = 0.3$$

$$[Cl_2]_{eq} = 1 + x = \boxed{1.3M}$$

$$[HCl]_{eq} = 1 - 2(0.3) = \boxed{0.4M}$$

$$K_{eq} = \frac{[HCl]^2}{[H_2][Cl_2]} = \frac{(0.4)^2}{(1.3)(1.3)} = \boxed{0.094}$$

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**Roxy Wilson, Theodore L. Brown, Harold
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