# Principles of combustion

- Combustion refers to the rapid oxidation of fuel accompanied by the production of heat, or heat and light.
- Complete combustion (complete oxidation) of a fuel is possible only in the presence of an adequate supply of oxygen. (Air: 20.9% oxygen, 79% nitrogen and other gases).
- Rapid fuel oxidation results in large amounts of heat.
- Solid or liquid fuels must be changed to a gas before they will burn.
- Fuel gases will burn in their normal state if enough air is present.
- Nitrogen is considered to be a temperature reducing diluter.
- Nitrogen reduces combustion efficiency by absorbing heat from the combustion of fuels and diluting the flue gases.
- Nitrogen reduces the heat available for transfer through the heat exchanger surfaces.
- It also increases the volume of combustion by-products, which then have to travel through the heat exchanger and up the stack faster to allow the introduction of additional fuel-air mixture.
- Nitrogen also can combine with oxygen (particularly at high flame temperatures) to produce oxides of nitrogen (NO<sub>2</sub>), which are toxic pollutants.
- Carbon may also combine with oxygen to form carbon monoxide, which results in the release of a smaller quantity of heat. Carbon burned to CO<sub>2</sub> will produce more heat per unit of fuel than when CO or smoke are produced.

# **Principles Of Combustion**

**David Ames Wells** 

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**Principles of Combustion** Kenneth K. Kuo, 1986-05-08 This comprehensive text covers principles and applications with an emphasis on the theoretical modeling of combustion Addresses chemical thermodynamics and kinetics conservation equations for multi component reacting flows deflagration and detonation waves premixed laminar flames spray combustion of fuel droplets ignition and related topics Many examples are included to demonstrate the application of theory Emphasizes **Principles of Combustion** Allan T. Kirkpatrick, Kenneth K. Kuo, 2024-11-27 the use of digital computers for solutions The new edition of a classic textbook on combustion principles and processes covering the latest developments in fuels and applications in a student friendly format Principles of Combustion provides clear and authoritative coverage of chemically reacting flow systems Detailed and accessible chapters cover key combustion topics such as chemical kinetics reaction mechanisms laminar flames droplet evaporation and burning and turbulent reacting flows Numerous figures end of chapter problems extensive reference materials and examples of specific combustion applications are integrated throughout the text Newly revised and expanded Principles of Combustion makes it easier for students to absorb and master each concept covered by presenting content through smaller bite sized chapters Two entirely new chapters on turbulent reacting flows and solid fuel combustion are accompanied by additional coverage of low carbon fuels such as hydrogen natural gas and renewable fuels This new edition contains a wealth of new homework problems new application examples up to date references and access to a new companion website with MATLAB files that students can use to run different combustion cases Fully updated to meet the needs of today's students and instructors Principles of Combustion Provides problem solving techniques that draw from thermodynamics fluid mechanics and chemistry Addresses contemporary topics such as zero carbon combustion turbulent combustion and sustainable fuels Discusses the role of combustion emissions in climate change and the need for reducing reliance on carbon based fossil fuels Covers a wide range of combustion application areas including internal combustion engines industrial heating and materials processing Containing both introductory and advanced material on various combustion topics Principles of Combustion Third Edition is an essential textbook for upper level undergraduate and graduate courses on combustion combustion theory and combustion processes It is also a valuable reference for combustion engineers and scientists wanting to better understand a particular combustion problem

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