

CHEM 131 Study Guide:

Chapter 1

- ❖ Calculating density from mass and volume: $\text{Density} = \frac{\text{mass}}{\text{volume}}$
- ❖ Converting temperature from °C to K: $T(\text{in K}) = T(\text{in } ^\circ\text{C}) + 273.15$
- ❖ Converting temperature from K to °C: $T(\text{in } ^\circ\text{C}) = T(\text{in K}) - 273.15$
- ❖ Converting temperature from °C to °F: $T(\text{in } ^\circ\text{F}) = \frac{9}{5}T(\text{in } ^\circ\text{C}) + 32$
- ❖ Converting temperature from °F to °C: $T(\text{in } ^\circ\text{C}) = [T(\text{in } ^\circ\text{F}) - 32] \times \frac{5}{9}$

Chapter 2

- ❖ Finding the mass of an element in a given mass of compound:
$$\text{Mass of element in sample} = \text{mass of compound in sample} \times \frac{\text{mass of element in compound}}{\text{mass of compound}}$$
- ❖ Calculating the number of neutrons in an atom:
$$\text{Number of neutrons} = \text{mass number} - \text{atomic number}$$

or
$$N = A - Z$$
- ❖ Determining the molecular mass of a formula unit of a compound:
$$\text{Molecular mass} = \text{sum of atomic masses}$$

Chapter 3

- ❖ Number of entities in one mole: 1 mol contains 6.022×10^{23} entities (to 4 sf)
- ❖ Converting amount (mol) to mass (g) using M :
$$\text{Mass (g)} = \text{amount (mol)} \times \frac{\text{no. of grams}}{1 \text{ mol}}$$
- ❖ Converting mass (g) to amount (mol) using $1/M$:
$$\text{Amount (mol)} = \text{mass (g)} \times \frac{1 \text{ mol}}{\text{no. of grams}}$$
- ❖ Converting amount (mol) to number of entities :
$$\text{No. of entities} = \text{amount (mol)} \times \frac{6.022 \times 10^{23} \text{ entities}}{1 \text{ mol}}$$
- ❖ Converting number of entities to amount (mol):
$$\text{Amount (mol)} = \text{no. of entities} \times \frac{1 \text{ mol}}{6.022 \times 10^{23} \text{ entities}}$$

$$\text{Mass \% of element X} = \frac{\text{moles of X in formula} \times \text{molar mass of X (g/mol)}}{\text{mass (g) of 1 mol of compound}} \times 100$$
- ❖ Calculating mass %:
- ❖ Finding the mass of an element in any mass of compound:
$$\text{Mass of element} = \text{mass of compound} \times \frac{\text{mass of element in 1 mol of compound}}{\text{mass of 1 mol of compound}}$$
- ❖ Calculating percent yield:
$$\% \text{ yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$$

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