

CHEM 1211 Sep 28

Friday, September 30, 2022 9:37 AM

5c) Gas Stoichiometry

Gas Stoichiometry

- Much like we could use molarity and volume to find the moles of an aqueous compound, we can use pressure, volume, and temperature to solve for the moles of a gas to use in stoichiometry.

5c) Kinetic Molecular Theory

Postulates of Kinetic Molecular Theory

- The volume of individual gas particles are negligible (point particles)
- Particles of gas are in constant motion
 - The collision of the gas particles with the walls of the container causes pressure
- The particles of gases don't attract or repel each other-- elastic collisions
- The average kinetic energy of a gas is directly proportional to the temperature (Kelvin) of a gas.

How does this explain Dalton's Law?

$$P_{\text{total}} = P_1 + P_2 + P_3$$

Velocity of a Gas

- The kelvin temperature of a gas is proportional to the kinetic energy

$$(KE)_{\text{avg}} = \frac{3}{2} RT$$

- Root mean square velocity is a special average of the velocities, that keeps opposite velocities from cancelling out.

$$U_{\text{rms}} = \text{sqrt} \left(\frac{3RT}{M} \right)$$

- Where R= 8.314 J/K*mol and M is the mass in kilograms of a mole of the gas (molar mass divided by 1000)

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Table of Contents Pkgacp Chem 1211

1. Understanding the eBook Pkgacp Chem 1211
 - The Rise of Digital Reading Pkgacp Chem 1211
 - Advantages of eBooks Over Traditional Books
2. Identifying Pkgacp Chem 1211
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pkgacp Chem 1211
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pkgacp Chem 1211
 - Personalized Recommendations

- Pkgacp Chem 1211 User Reviews and Ratings
- Pkgacp Chem 1211 and Bestseller Lists
- 5. Accessing Pkgacp Chem 1211 Free and Paid eBooks
 - Pkgacp Chem 1211 Public Domain eBooks
 - Pkgacp Chem 1211 eBook Subscription Services
 - Pkgacp Chem 1211 Budget-Friendly Options
- 6. Navigating Pkgacp Chem 1211 eBook Formats
 - ePub, PDF, MOBI, and More
 - Pkgacp Chem 1211 Compatibility with Devices
 - Pkgacp Chem 1211 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Pkgacp Chem 1211
 - Highlighting and Note-Taking Pkgacp Chem 1211
 - Interactive Elements Pkgacp Chem 1211
- 8. Staying Engaged with Pkgacp Chem 1211
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pkgacp Chem 1211
- 9. Balancing eBooks and Physical Books Pkgacp Chem 1211
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pkgacp Chem 1211
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Pkgacp Chem 1211
 - Setting Reading Goals Pkgacp Chem 1211
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Pkgacp Chem 1211
 - Fact-Checking eBook Content of Pkgacp Chem 1211

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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