CHEM 1120-03 - General Chemistry 2

Exam Reference Page (to be provided during exams)

Universal Information:

 $\Delta S_{univ} = \Delta S_{sys} + \Delta S_{surr}$

 $\Delta S^{o} = \sum mS^{o}_{necoloris} - \sum nS^{o}_{reactants}$

$$T(K) = T(^{*}C) + 273.15$$
 For the eq: $ax^{2} + bx + c = 0$, $x = \frac{-b \pm \sqrt{b^{2} - 4ac}}{2a}$
 $R = 8.314 \frac{J}{mol \cdot K} = 0.0821 \frac{L \cdot atm}{mol \cdot K}$ Faraday's const. = $9.6485 \times 10^{4} \frac{C}{mol}$
 $1 \text{ V} = 1 \text{ J/C}$ and $1 \text{ A} = 1 \text{ C/s}$
 $1 \text{ Ci} = 3.70 \times 10^{10} \text{ Bg} = 3.70 \times 10^{30} \text{ d/s}$ $1.000 \text{ amu} = 9.321 \times 10^{8} \text{ eV} = 932.91 \text{ MeV}$

Useful equations and relationships:

$S_{gas} = k_M \times P_{gas}$		Zero Order	First Order	Second Order
$\Delta P = i(X_{solute} \times P^{*}_{solvent})$	Rate law Units for k	rate = & mod/L-x	rate = k[A] 1/s	rate = $k[A]^2$ L/mol·s
$\Delta T_b = i(K_b m)$	Integrated rate law in straight-line form	$[A]_{r} =$ $-\epsilon_{r} + [A]_{0}$	$ln [A]_i =$ $-kr + ln [A]_0$	$1/[A]_r = $ $4r + 1/[A]_0$
$\Delta T_f = i(K_f m)$	Plot for straight line Slope, y intercept	[AL vs. time -k, [Alo	In [A], vs. time -k, in [A]o	L/[A], vs. time & L/[A]o
$\Pi = i(MRT)$	Half-life	(A) ₀ /28	(lm 2)/k	1/8[A]o

 $\Delta H^{o} = \sum m\Delta H_{t,products}^{o} - \sum n\Delta H_{t,products}^{o}$

 $\Delta G^{\circ} = \sum m\Delta G_{t,prostucts}^{\circ} - \sum n\Delta G_{t,prostucts}^{\circ}$

$$\Delta G^\circ = -nFE_{\rm cell}^\circ \qquad \qquad E_{\rm cell}^\circ = E_{\rm cathode}^\circ - E_{\rm anode}^\circ \qquad \qquad E_{\rm cell} = E_{\rm cell}^\circ - \frac{\kappa T}{nF} \ln Q$$

$$E_{\rm cell}^\circ = \frac{\kappa T}{nF} \ln K \qquad \qquad E_{\rm cell} = E_{\rm cell}^\circ - \frac{0.0592\,V}{n} \log Q$$

$$\ln \frac{N_\ell}{N_0} = \ln \frac{A_\ell}{A_0} = \ln \frac{m_\ell}{m_0} = -kt$$

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