

## CHEM 101 Lab - Types of Chemical Reactions

**General Discussions:** A chemical reaction produces one or more new substances. The reaction is indicated from one or more of the observations which may be: *formation of solid precipitate, energy change, evolution of gas, production of color or odor, and change in the pH of the solution.*

When we say that a solid precipitate forms, that means one of the products of the reaction is a solid, insoluble in water. That is, its solubility product is high, therefore the product formed precipitates out from the solution as an insoluble solid (or precipitate) which settles down at the bottom of the reaction container. For example, if you add hydrochloric acid, which is an acid to barium hydroxide, a base, you should expect the formation of barium chloride, which is a salt along with water. However, you won't see any solid precipitate of barium chloride because it is soluble in water. On the other hand, if you mix sulfuric acid with barium hydroxide you will immediately see a white precipitate formed. This white solid is barium sulfate which is insoluble in water.

Many reactions produce gaseous products, such as hydrogen, oxygen or carbon dioxide. In such reactions you will see the gas bubbles evolving out. Additionally,  $O_2$  gas can be tested by exposing it to a glowing splinter. It would set alight the glowing splinter. Similarly, when  $H_2$  gas is exposed to a burning match stick it would make a popping noise. Many reactions proceed with energy change. That is heat is absorbed or released in the reaction. If you have to heat the mixture to carry out the reaction, that means heat is absorbed and the reaction is endothermic and the energy change process is designated by a positive sign (+). On the other hand, if heat is released in the reaction, such process is termed as exothermic and the magnitude of energy change is designated by a negative sign (—). In an exothermic reaction, you may feel your test tube becoming warmer or heated due to heat release. The opposite happens in the endothermic reactions. Other form of energy may be involved in the reaction, such as light (examples are photochemical reactions). Acid-base reactions are indicated by change in the pH of the solution.

**Chemical Equations:** Every chemical reaction can be represented by a chemical equation that must be balanced. A balanced chemical equation indicates what are the substances reacting and what the products of the reactions are. Additionally, it tells the quantities of substances reacting and formed in moles, grams, and the number of molecules and atoms. Often it tells the conditions under which the reactions carried out, such as temperature, pressure and the catalyst if any is used; and also, the physical states of the substances reacting and produced, such as, solid (s), liquid (l), gas (g), or aqueous solution (aq). A few examples are presented below in this discussion.

**Types of reactions:** There are broadly five types of reactions: combination, decomposition, single replacement (or single displacement), double replacement (or double displacement or

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**Ian Pickup**

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