#### Chapter 13: Colligative Properties of Solutions

#### Sections to read before coming to class:

13-1, 13-2, 13-3, 13-4, 13-5, and 13-6

#### I. Objectives

By the end of this chapter, you should be able to describe:

- A. Definitions
- B. Concentration units
  - Molar concentration and molal concentration
  - 2. Mole fraction and mole percent
  - Weight fraction and weight percent.
- C. The Solution Process
  - 1. Heats of solution
  - "Like dissolves like".
  - Factors affecting solubility.
- D. Colligative properties.
  - 1. Vapor pressure and Raoult's Law
  - Boiling point elevation and freezing point depression.
  - Molar masses from colligative properties
  - 4. Osmotic pressure
- II. Definitions
- A Solution: homogenous mixture of a solute dissolved in a solvent mixture).
  - B. Saturated solution: Dissolved solute is in equilibrium with solid solute solute (aq) Solid (
  - amount of a solute regid C. Solubility: to make a saturated solution

(temperature dependent)

D. Unsaturated solution: Solution has less solute than required to make a saturated. solution.

# Pkgacp Concordia U Chem 1401

**Brooks/Cole** 

# Pkgacp Concordia U Chem 1401:

**Acp Chem 241 - Concordia University** Brooks/Cole,2013-08-19 Liotta,2008

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- 1. The book is structured into several chapters, namely:
  - Chapter 1: Introduction to Pkgacp Concordia U Chem 1401
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  - ∘ Chapter 3: Pkgacp Concordia U Chem 1401 in Everyday Life
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  - ∘ Chapter 5: Conclusion
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- 4. In chapter 3, the author will examine the practical applications of Pkgacp Concordia U Chem 1401 in daily life. This chapter will showcase real-world examples of how Pkgacp Concordia U Chem 1401 can be effectively utilized in everyday scenarios.
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- 6. In chapter 5, this book will draw a conclusion about Pkgacp Concordia U Chem 1401. This chapter will summarize the key points that have been discussed throughout the book.
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