

Chapter 4: Solutions Calculations and Understanding Precipitation and Acid/Base Reactions*Sections to read before coming to class:*

4-1, 4-2, 4-3, and 4-4

I. Objectives**A. Solutions**

1. Properties of solutes in aqueous solutions
2. Solubilities of ionic compounds in water

B. Reactions in aqueous solutions

1. Precipitation reactions
2. Acids and bases

C. Solution stoichiometry

1. Molar concentrations
2. Preparing solutions
3. Determining pH, a measure of acidity
4. Stoichiometry of reactions in solution
5. Titrations

II. Solutions

Hmmm....why study salt and water?

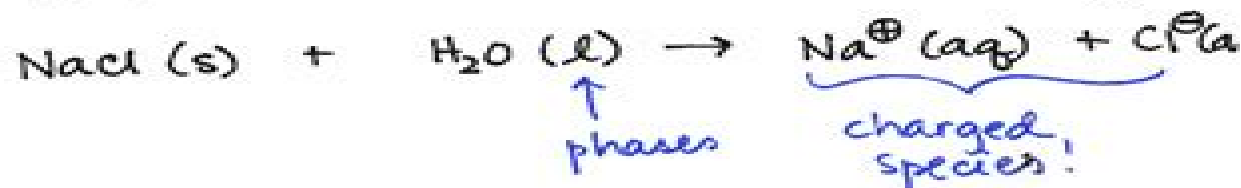
Biology

1. Water and salt are essential for living systems to function.
2. **Aqueous solutions** are chemical systems in which water plays a role.
3. **Solutions** are homogeneous mixtures of a solute dissolved in a solvent.
4. Focus first on ionic compounds (strong electrolytes) in water.

metal + non-metal bonded together**A. Properties of solutes in aqueous solutions****1. Electrolytes vs. nonelectrolytes**

- a. A special property of water allows for electrolytes (ionic compounds) to:

dissociate and conduct electricity



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