

CHEM 1211 Sep 28

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5c) Gas Stoichiometry

Gas Stoichiometry

- Much like we could use molarity and volume to find the moles of an aqueous compound, we can use pressure, volume, and temperature to solve for the moles of a gas to use in stoichiometry.

5c) Kinetic Molecular Theory

Postulates of Kinetic Molecular Theory

- The volume of individual gas particles are negligible (point particles)
- Particles of gas are in constant motion
 - The collision of the gas particles with the walls of the container causes pressure
- The particles of gases don't attract or repel each other-- elastic collisions
- The average kinetic energy of a gas is directly proportional to the temperature (Kelvin) of a gas.

How does this explain Dalton's Law?

$$P_{\text{total}} = P_1 + P_2 + P_3$$

Velocity of a Gas

- The kelvin temperature of a gas is proportional to the kinetic energy

$$(KE)_{\text{avg}} = \frac{3}{2} RT$$

- Root mean square velocity is a special average of the velocities, that keeps opposite velocities from cancelling out.

$$U_{\text{rms}} = \text{sqrt} \left(\frac{3RT}{M} \right)$$

- Where R= 8.314 J/K*mol and M is the mass in kilograms of a mole of the gas (molar mass divided by 1000)

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