

CHEM 101 Tutorial L6

Chemical Equilibrium

- Based on these descriptions, write a balanced equation and the corresponding K_c expression for each reversible reaction.
 - Carbonyl fluoride, $\text{COF}_2(\text{g})$, decomposes into gaseous carbon dioxide and gaseous carbon tetrafluoride.
 - Peroxodisulfate ion, $\text{S}_2\text{O}_8^{2-}$, oxidises iron(II) ion to iron(III) ion in aqueous solution and is itself reduced to sulfate ion.
- Write equilibrium constant expressions, K_c , for the reactions.
 - $2 \text{NO}(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2 \text{NO}_2(\text{g})$
 - $\text{Mg}(\text{OH})_2(\text{s}) + \text{CO}_3^{2-}(\text{aq}) \rightleftharpoons \text{MgCO}_3(\text{s}) + 2 \text{OH}^-(\text{aq})$
- Write equilibrium constant expressions, K_p , for the reactions.
$$\text{CS}_2(\text{g}) + 4 \text{H}_2(\text{g}) \rightleftharpoons \text{CH}_4(\text{g}) + 2 \text{H}_2\text{S}(\text{g})$$
- Calculate the value of K_c from the K_p value given.
$$2 \text{NO}(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2 \text{NO}_2(\text{g}) \quad K_p = 1.48 \times 10^4 \text{ at } 184^\circ\text{C}$$
- A mixture of 1.00 g H_2 and 1.06 g H_2S in a 0.500 L flask comes to equilibrium at 1670 K: $2 \text{H}_2(\text{g}) + \text{S}_2(\text{g}) \rightleftharpoons 2 \text{H}_2\text{S}(\text{g})$. The equilibrium amount of $\text{S}_2(\text{g})$ found is 8.00×10^{-6} mol. Calculate the value of K_p at 1670 K for pressures expressed in atmospheres.
- Compute K_c for the reaction $\text{N}_2(\text{g}) + \text{O}_2(\text{g}) + \text{Cl}_2(\text{g}) \rightleftharpoons 2 \text{NOCl}(\text{g})$, given the following data at 298 K.



- Equilibrium is established at 1000 K where $K_c = 281$ for the reaction $2 \text{SO}_2(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2 \text{SO}_3(\text{g})$. The equilibrium amount of $\text{O}_2(\text{g})$ in a 0.185 L flask is 0.00247 mol. Predict the ratio of $[\text{SO}_2]$ to $[\text{SO}_3]$ in this equilibrium mixture.
- An equilibrium mixture at 1000 K contains 0.276 mol H_2 , 0.276 mol CO_2 , 0.224 mol CO , and 0.224 mol H_2O .
$$\text{CO}_2(\text{g}) + \text{H}_2(\text{g}) \rightleftharpoons \text{CO}(\text{g}) + \text{H}_2\text{O}(\text{g})$$
 - Show that for this reaction, K_c is independent of the reaction volume, V .
 - Calculate the value of K_c and K_p .

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Table of Contents Pkgacp Chem 101

1. Understanding the eBook Pkgacp Chem 101
 - The Rise of Digital Reading Pkgacp Chem 101
 - Advantages of eBooks Over Traditional Books
2. Identifying Pkgacp Chem 101
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pkgacp Chem 101
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pkgacp Chem 101
 - Personalized Recommendations
 - Pkgacp Chem 101 User Reviews and Ratings

- Pkgacp Chem 101 and Bestseller Lists
- 5. Accessing Pkgacp Chem 101 Free and Paid eBooks
 - Pkgacp Chem 101 Public Domain eBooks
 - Pkgacp Chem 101 eBook Subscription Services
 - Pkgacp Chem 101 Budget-Friendly Options
- 6. Navigating Pkgacp Chem 101 eBook Formats
 - ePub, PDF, MOBI, and More
 - Pkgacp Chem 101 Compatibility with Devices
 - Pkgacp Chem 101 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Pkgacp Chem 101
 - Highlighting and Note-Taking Pkgacp Chem 101
 - Interactive Elements Pkgacp Chem 101
- 8. Staying Engaged with Pkgacp Chem 101
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pkgacp Chem 101
- 9. Balancing eBooks and Physical Books Pkgacp Chem 101
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pkgacp Chem 101
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Pkgacp Chem 101
 - Setting Reading Goals Pkgacp Chem 101
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Pkgacp Chem 101
 - Fact-Checking eBook Content of Pkgacp Chem 101
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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