

CHEM 101 Tutorial L6

Chemical Equilibrium

- Based on these descriptions, write a balanced equation and the corresponding K_c expression for each reversible reaction.
 - Carbonyl fluoride, $\text{COF}_2(\text{g})$, decomposes into gaseous carbon dioxide and gaseous carbon tetrafluoride.
 - Peroxodisulfate ion, $\text{S}_2\text{O}_8^{2-}$, oxidises iron(II) ion to iron(III) ion in aqueous solution and is itself reduced to sulfate ion.
- Write equilibrium constant expressions, K_c , for the reactions.
 - $2 \text{NO}(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2 \text{NO}_2(\text{g})$
 - $\text{Mg}(\text{OH})_2(\text{s}) + \text{CO}_3^{2-}(\text{aq}) \rightleftharpoons \text{MgCO}_3(\text{s}) + 2 \text{OH}^-(\text{aq})$
- Write equilibrium constant expressions, K_p , for the reactions.
$$\text{CS}_2(\text{g}) + 4 \text{H}_2(\text{g}) \rightleftharpoons \text{CH}_4(\text{g}) + 2 \text{H}_2\text{S}(\text{g})$$
- Calculate the value of K_c from the K_p value given.
$$2 \text{NO}(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2 \text{NO}_2(\text{g}) \quad K_p = 1.48 \times 10^4 \text{ at } 184^\circ\text{C}$$
- A mixture of 1.00 g H_2 and 1.06 g H_2S in a 0.500 L flask comes to equilibrium at 1670 K: $2 \text{H}_2(\text{g}) + \text{S}_2(\text{g}) \rightleftharpoons 2 \text{H}_2\text{S}(\text{g})$. The equilibrium amount of $\text{S}_2(\text{g})$ found is 8.00×10^{-6} mol. Calculate the value of K_p at 1670 K for pressures expressed in atmospheres.
- Compute K_c for the reaction $\text{N}_2(\text{g}) + \text{O}_2(\text{g}) + \text{Cl}_2(\text{g}) \rightleftharpoons 2 \text{NOCl}(\text{g})$, given the following data at 298 K.



- Equilibrium is established at 1000 K where $K_c = 281$ for the reaction $2 \text{SO}_2(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2 \text{SO}_3(\text{g})$. The equilibrium amount of $\text{O}_2(\text{g})$ in a 0.185 L flask is 0.00247 mol. Predict the ratio of $[\text{SO}_2]$ to $[\text{SO}_3]$ in this equilibrium mixture.
- An equilibrium mixture at 1000 K contains 0.276 mol H_2 , 0.276 mol CO_2 , 0.224 mol CO , and 0.224 mol H_2O .
$$\text{CO}_2(\text{g}) + \text{H}_2(\text{g}) \rightleftharpoons \text{CO}(\text{g}) + \text{H}_2\text{O}(\text{g})$$
 - Show that for this reaction, K_c is independent of the reaction volume, V .
 - Calculate the value of K_c and K_p .

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