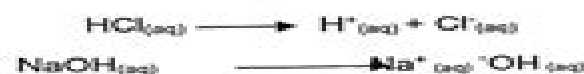


According to ARRHENIUS concept, an acid is a compound that release  $H^+$  ions in water, and base is a compound that release  $OH^-$  ions in water.

For example: HCl is an Arrhenius acid NaOH, is an Arrhenius base.



## Classes of Acid and Bases (Brønsted)

Strong acids/bases are strong electrolytes.

1. Strong acids/bases which are capable of donating/accepting protons.



2. Strong acids/bases which are capable of losing or gaining protons.



3. Strong/weak acids/bases (donates or loses their protons or accepts or releases protons from water) called strong/weak electrolytes.

For example:  $H_2SO_4, HCl, HNO_3$  water acts as a base by accepting protons from the acids.



Similarly, water acts as an acid while donating protons to electrolytes.



Stronger acid concentration of a solution is called pH. It is defined as the negative of base-10 logarithm of  $H^+$  concentration.

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