



**Figure 11.6** A pH meter is commonly used in the laboratory to determine the pH of a solution. Although many pH meters have scales marked with values from 1 to 14, pH values can, in fact, be less than 1 and greater than 14. The science behind the pH meter will be discussed in Chapter 13.

In the laboratory, the pH of a solution is measured with a pH meter (Figure 11.6), which measures the activity of the  $\text{H}^+$  ion electrochemically. (The construction of a pH meter is discussed in more detail in Chapter 13). Table 11.1 lists the pH values of a number of common fluids at  $25^\circ\text{C}$ . As you can see, the pH of body fluids varies greatly, depending on location and function. The low pH (high acidity) of gastric juices facilitates digestion whereas the higher pH of blood is necessary for the transport of oxygen.

A pOH scale analogous to the pH scale can be devised using the negative logarithm (base 10) of the hydroxide ion activity of a solution. Thus, we define pOH as

$$\text{pOH} = -\log_{10} a_{\text{OH}^-} \quad (11.6)$$

As in the case of pH, the pOH for sufficiently dilute aqueous solutions can be written as

$$\text{pOH} \approx -\log_{10} [\text{OH}^-] \quad (11.7)$$

Now consider again the ion-product constant for water at  $25^\circ\text{C}$ :

$$[\text{H}^+][\text{OH}^-] = K_w = 1.01 \times 10^{-14}$$

Taking the negative logarithm of both sides, we obtain

$$\begin{aligned} -\log_{10} [\text{H}^+] - \log_{10} [\text{OH}^-] &= -\log_{10} (1.01 \times 10^{-14}) \\ &= 14.00 \end{aligned}$$

From the definitions of pH and pOH we obtain

$$\text{pH} + \text{pOH} = 14.00 \quad (11.8)$$

at  $25^\circ\text{C}$ . Equation 11.8 provides us with another way to express the relationship between the  $\text{H}^+$  ion concentration and the  $\text{OH}^-$  ion concentration.

Examples 11.4 to 11.6 illustrate calculations involving pH at  $25^\circ\text{C}$ .

### Example 11.4

The concentration of  $\text{H}^+$  ions in a bottle of table wine was  $3.2 \times 10^{-4} \text{ M}$  right after the cork was removed. Only half of the wine was consumed. The other half, after it

—Continued

**Table 11.1**  
The pH Values of Some Common Fluids

Sample	pH Value
Gastric juice in the stomach	1.0–2.0
Lemon juice	2.4
Vinegar	3.0
Grapefruit juice	3.2
Orange juice	3.5
Urine	4.8–7.5
Water exposed to air*	5.5
Saliva	6.4–6.9
Milk	6.5
Pure water	7.0
Blood	7.35–7.45
Tears	7.4
Milk of magnesia	10.6
Household ammonia	11.5

\*Water exposed to air for a long period of time absorbs atmospheric  $\text{CO}_2$  to form carbonic acid,  $\text{H}_2\text{CO}_3$ .

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