

Chemistry 1100

Lecture 1

Energy of Electromagnetic Radiation: $E = h\nu$, E = Energy, h = Planck's Constant, ν = frequency

Low energy = low frequency = long wavelength = low amplitude for light

High energy = high frequency = short wavelength = high amplitude for light

Light has wave-like properties: $c = \nu\lambda$, c = speed of light, ν = frequency, λ = wavelength

Critical relationship between the energy of a wave and the wavelength of the wave!!!!

$$E = \frac{hc}{\lambda}$$

- Light particles: photons
- Photon properties include energy, velocity, momentum, zero resting mass (basically if they don't move they don't exist), and no electrical charge.
- Photons travel with the speed of light within a vacuum.
- Photons can be seen as tiny packages of electromagnetic energy, this is what Einsteins Planck equation represents: $E = h\nu$.
- Electromagnetic radiation (light) exhibits wave properties and particulate properties.



Light as wave

- Light has frequency, wavelength, and amplitude

$$\nu = \frac{c}{\lambda}$$

- c = velocity of light ($3 \times 10^8 \text{ m s}^{-1}$)
- ν = frequency (s^{-1})
- λ = wavelength (m)

Light as particle

- Light has **particle** like properties – “quanta” (or photons)

$$E = h\nu$$

- h = Planck's constant = $6.626 \times 10^{-34} \text{ J s}$
- A quantum of energy = $h\nu$
- In atoms and molecules, **energy** is given out (or accepted) in whole number packets

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Daniela Niemeyer



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Table of Contents Pkg Acp Chemistry 1100

1. Understanding the eBook Pkg Acp Chemistry 1100
 - The Rise of Digital Reading Pkg Acp Chemistry 1100
 - Advantages of eBooks Over Traditional Books
2. Identifying Pkg Acp Chemistry 1100
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pkg Acp Chemistry 1100
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pkg Acp Chemistry 1100
 - Personalized Recommendations

- Pkg Acp Chemistry 1100 User Reviews and Ratings
- Pkg Acp Chemistry 1100 and Bestseller Lists
- 5. Accessing Pkg Acp Chemistry 1100 Free and Paid eBooks
 - Pkg Acp Chemistry 1100 Public Domain eBooks
 - Pkg Acp Chemistry 1100 eBook Subscription Services
 - Pkg Acp Chemistry 1100 Budget-Friendly Options
- 6. Navigating Pkg Acp Chemistry 1100 eBook Formats
 - ePub, PDF, MOBI, and More
 - Pkg Acp Chemistry 1100 Compatibility with Devices
 - Pkg Acp Chemistry 1100 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Pkg Acp Chemistry 1100
 - Highlighting and Note-Taking Pkg Acp Chemistry 1100
 - Interactive Elements Pkg Acp Chemistry 1100
- 8. Staying Engaged with Pkg Acp Chemistry 1100
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pkg Acp Chemistry 1100
- 9. Balancing eBooks and Physical Books Pkg Acp Chemistry 1100
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pkg Acp Chemistry 1100
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Pkg Acp Chemistry 1100
 - Setting Reading Goals Pkg Acp Chemistry 1100
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Pkg Acp Chemistry 1100
 - Fact-Checking eBook Content of Pkg Acp Chemistry 1100

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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