

Special Considerations:

Hydrochloric acid can cause burns to the skin or eyes. If you get hydrochloric acid on your skin or in your eyes, wash the affected area with lots of water.

Boric acid is toxic to humans and can irritate the skin. Wash your hands well after completing this experiment.

Procedure:

- 1) Your instructor will have set up five saturated solutions of sodium tetraborate at five different temperatures. Thermometers will be in each solution so you can record the temperature of each just before you use it.
- 2) Place a 150 mL beaker containing about 100 mL of deionized water onto the hot plate and allow it to heat.
- 3) Set up a buret, rinse and fill it with standardized hydrochloric acid solution that has been provided. Record the molarity of the acid on your data sheet.
- 4) You'll titrate each of the five saturated solutions your instructor set up in turn.
- 5) Record the temperature of the mixture on your data sheet and then immediately decant 5.0 mL of the saturated solution into a graduated cylinder. Be sure not to add any solid from the bottom of the mixture to your graduated cylinder. You may see solid begin to form when some of the warmer solutions begin to cool in the graduated cylinder. This is normal.
- 6) IMMEDIATELY and quantitatively transfer the contents of the graduated cylinder to a 250 mL Erlenmeyer flask. Use the hot deionized water you are heating to wash the graduated cylinder. You'll likely see precipitate form when the hotter saturated solutions come into contact with the cold graduated cylinder. All the precipitate needs to be washed into your 250 mL Erlenmeyer flask. Make sure that the transfer of the borax into the Erlenmeyer flask is quantitative.
- 7) Add enough deionized water to the Erlenmeyer flask to make the total volume about 75 mL.
- 8) Add 3 drops of methyl orange indicator to the contents of the flask and titrate with HCl. The equivalence point is reached when the indicator turns from yellow to red. Record starting and ending buret readings on your data sheet.
- 9) Repeat steps 5-8 at least four more times making sure you do each titration using saturated borax solution at a temperature that is at least 5°C less than the previous trial. Make sure you record all your data on the data sheet.
- 10) Calculate the K_{sp} value for your first trial.
- 11) Enter all of your data onto a spreadsheet. Make a plot of $\ln K$ vs. $\frac{1}{T}$ and calculate ΔH and ΔS from the plot.

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