# MISSED THIS? Read Section 16.3. You can click on the Review had to access the section in your exert.

The following reaction has an equilibrium constant of  $K_{\rm P}=2.26\times 10^4$  at 298  ${
m K}_{\odot}$ 

$$\mathrm{CO}(g) + 2\mathrm{H}_2(g) \Longrightarrow \mathrm{CH}_2\mathrm{OH}(g)$$

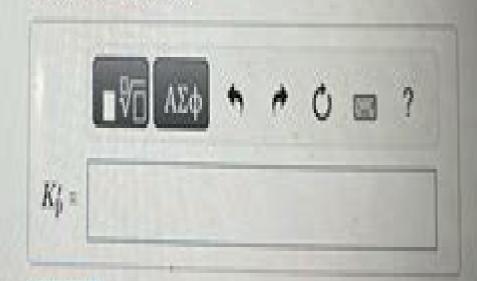
# ▼ Part A

Calculate  $K_{\rm P}^{\rm r}$  for the following reaction:

$$\mathrm{CH_2OH}(\mathrm{g}) \rightleftharpoons \mathrm{CO}(\mathrm{g}) + 2\mathrm{H_2}(\mathrm{g})$$

Express your answer to three significant figures.

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**Boyles** 

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