

**II. PREPARATION OF ACETYL SALICYLIC ACID (ASPIRIN):**

1. Complete the table below by entering appropriate values into the empty boxes. Show calculations for all entries clearly below.

Reactants	Chemical Formula ( $C_6H_7O_2$ )	MW (g/mol) $\pm 0.01$ g	Grams of Reactant	Moles of Reactant	Moles of Aspirin <sup>1</sup>	Limiting Reagent (put X)	Melting Range
Salicylic Acid							
Acetic Anhydride							

- 1 The moles of aspirin from each reactant if that reactant were the limiting reagent (show calculations below).
- 2 Calculate the grams of acetic anhydride from the density = \_\_\_\_\_ g/mL and the volume you used: \_\_\_\_\_ mL. Show work below:

Show all calculations for the above table, below making sure to include units:

Molecular Weight: Acetic Anhydride  
(use atomic weights  $\pm 0.01$  g)

Molecular Weight: Salicylic Acid  
(use atomic weights  $\pm 0.01$  g)

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