

CHAPTER 12.1: Chemical Reactions

Introduction One of the most important laboratory tools available to chemists after conducting experiments and observing chemical reactions is stoichiometry.

$$\text{Reactants} \rightarrow \text{Products}$$

(Molar Masses)
(Molar Masses)
(Molar Masses)

STOICHIOMETRY Many experiments that chemists carry out involve a chemical reaction. Many factors can affect how fast a reaction takes place and the amounts of products and reactants involved. But without some basic knowledge of stoichiometry, chemists cannot design experiments that lead to useful information. In this chapter, we will learn how to use stoichiometry to predict the amounts of products and reactants involved in a chemical reaction. We will also learn how to use stoichiometry to predict the amounts of products and reactants involved in a chemical reaction.

Stoichiometric Reactions $A + B \rightarrow C + D$

In a stoichiometric reaction, the reactants combine in a fixed ratio to form a single product. The most common example of a stoichiometric reaction is the reaction of hydrogen and oxygen to produce a single product, water. The reaction is shown below:



In this reaction, the reactants are hydrogen and oxygen, and the product is water. The reaction is shown below:



When this reaction is carried out in a laboratory, you will observe that hydrogen and oxygen combine to form water. The reaction is shown below:

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