

Acid Base Titration

Name _____

Section _____

Partner _____

Instructor _____

Grams of NaOH needed to make 500 mL of 0.3 M NaOH solution 6.00g
 (Show calculation and give correct significant figures) (2 pts.)

(Unless otherwise noted, all blanks are worth 2 pts.)

Standardization of NaOH

	Trial 1**	Trial 2	Trial 3
Mass of $\text{KC}_8\text{H}_5\text{O}_4$	<u>.514g</u>	<u>.506g</u>	<u>.513g</u>
Initial Buret Reading NaOH	<u>0 mL</u>	<u>8.89 mL</u>	<u>17.61 mL</u>
Final Buret Reading NaOH	<u>8.89 mL</u>	<u>17.61 mL</u>	<u>26.55 mL</u>
Total Volume NaOH added	<u>8.89 mL</u>	<u>8.72 mL</u>	<u>8.94 mL</u>
NaOH Molarity	_____	_____	_____ (4 pts each)

Average Molarity of NaOH _____

**All calculations Trial #1 (4 pts.)

$$\begin{array}{r}
 17.61 \\
 - 8.89 \\
 \hline
 8.72
 \end{array}
 \qquad
 \begin{array}{r}
 26.55 \\
 - 17.61 \\
 \hline
 8.94
 \end{array}$$

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