

CHEM 1450 Experiment 7 Post-Lab
Acid-Base Equilibria, pH, and Buffers

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Date: 4/9/24

Lab Partner: Amber Doyle

Lab Section: 126

Lab TA: Tom

Experimental Data, Observations, and Notebook Pages (1.0 pt)

- Attach copies of your lab notebook pages to this post-lab before turning it in.
- Neatly copy your collected experimental data from the laboratory notebook into the summary below:

A. Hydrochloric Acid Solutions

- | | | | |
|--------------------------------|----------------|--------------------------------|-------------|
| 1. Original concentration HCl | <u>.10</u> M | 2. pH of original HCl solution | <u>1.31</u> |
| 3. Volume HCl used in dilution | <u>5.01</u> mL | 4. pH of diluted HCl solution | <u>1.99</u> |

B. Acetic Acid Solutions

- | | | | |
|--|----------------|---|-------------|
| 5. Original conc. $\text{CH}_3\text{CO}_2\text{H}$ | <u>.10</u> M | 6. pH of original $\text{CH}_3\text{CO}_2\text{H}$ soln | <u>2.90</u> |
| 7. Volume $\text{CH}_3\text{CO}_2\text{H}$ used | <u>5.03</u> mL | 8. pH of diluted $\text{CH}_3\text{CO}_2\text{H}$ soln | <u>3.36</u> |

C. Buffer Solutions

- | | |
|---|--|
| 9. Mass of $\text{CH}_3\text{CO}_2\text{Na}$ used | <u>4.116</u> g |
| 10. Original concentration $\text{CH}_3\text{CO}_2\text{H}$ | <u>6.00</u> M |
| 11. Volume of $\text{CH}_3\text{CO}_2\text{H}$ used | <u>8.48</u> mL |
| 12. Volume of distilled water used | <u>91.52</u> mL |
| 13. Total volume of buffer solution | <u>100.00</u> mL |
| 14. Volume of 6 M HCl used for addition to beaker A: | <u>1.04</u> mL; beaker B: <u>1.04</u> mL |
| 15. Volume of 6 M NaOH used for addition to beaker C: | <u>1.02</u> mL; beaker D: <u>1.04</u> mL |

	<i>Initial pH of Solutions</i>	<i>pH of Solutions after additions</i>
16. Beaker A	<u>6.40</u>	<u>9.0</u>
17. Beaker B	<u>4.56</u>	<u>4.00</u>
18. Beaker C	<u>6.22</u>	<u>13.42</u>
19. Beaker D	<u>4.57</u>	<u>5.02</u>

Pkgacp Uca Chem 1450

Sebastian Brünink



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Table of Contents Pkgacp Uca Chem 1450

1. Understanding the eBook Pkgacp Uca Chem 1450
 - The Rise of Digital Reading Pkgacp Uca Chem 1450
 - Advantages of eBooks Over Traditional Books
2. Identifying Pkgacp Uca Chem 1450
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pkgacp Uca Chem 1450
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pkgacp Uca Chem 1450
 - Personalized Recommendations
 - Pkgacp Uca Chem 1450 User Reviews and Ratings
 - Pkgacp Uca Chem 1450 and Bestseller Lists

5. Accessing Pkgacp Uca Chem 1450 Free and Paid eBooks
 - Pkgacp Uca Chem 1450 Public Domain eBooks
 - Pkgacp Uca Chem 1450 eBook Subscription Services
 - Pkgacp Uca Chem 1450 Budget-Friendly Options
6. Navigating Pkgacp Uca Chem 1450 eBook Formats
 - ePub, PDF, MOBI, and More
 - Pkgacp Uca Chem 1450 Compatibility with Devices
 - Pkgacp Uca Chem 1450 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Pkgacp Uca Chem 1450
 - Highlighting and Note-Taking Pkgacp Uca Chem 1450
 - Interactive Elements Pkgacp Uca Chem 1450
8. Staying Engaged with Pkgacp Uca Chem 1450
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pkgacp Uca Chem 1450
9. Balancing eBooks and Physical Books Pkgacp Uca Chem 1450
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pkgacp Uca Chem 1450
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Pkgacp Uca Chem 1450
 - Setting Reading Goals Pkgacp Uca Chem 1450
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Pkgacp Uca Chem 1450
 - Fact-Checking eBook Content of Pkgacp Uca Chem 1450
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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