

Chem 105L4: Acid-Base Titration

Name: Jessica Pantoja Lab Section: _____ Date: 9/18/22

LAB PROCEDURE

PREPARATION

1. Gather your equipment: pipets (5 and 25 mL), rubber bulb, buret, ring stand, buret clamp, large beaker for waste, 125 and 250 mL E-flasks, 150, 250 and 400 mL beakers (labeled for vinegar, HCl and NaOH respectively), a watch glass to cover the NaOH beaker, and a small buret funnel. Phenolphthalein indicator solution (1% in alcohol) will be available, as well.
2. Clean and rinse the pipets and buret as discussed on pp. 2-3. Practice filling and transferring DI water with a pipet.
3. Obtain your reagents: 50 mL HCl (record the exact molarity of HCl: 0.955% M), and 100 mL NaOH (approximately — you can get more later if you need it). Obtain 20 mL of vinegar just before you need it—it is rather stinky!

TITRATION OF THE HCl SOLUTION

4. Rinse the walls of the buret completely with a small volume of the NaOH solution and then fill the buret with the same NaOH solution of unknown concentration.
5. Rinse the 20.00-mL (or 25.00-mL) pipet you are using with the standard HCl solution, then carefully pipette 20.00-mL (or 25.00-mL) of HCl into one of the E-flasks. (The flask must be clean, but need not be dry. Why?) Add 2-3 drops of the phenolphthalein indicator solution.
6. Read the buret (to two decimal places) and record this as the initial volume reading in Table 1.
7. Start the titration — carefully add NaOH to the HCl, swirling the flask, and observe any color change. As the equivalence point (endpoint) is being approached, the pink color will take longer to disappear, and you should add NaOH at a slower rate, until finally, you are adding NaOH dropwise. The exact endpoint is reached when a single drop of base is added and the pink color persists for at least 15 seconds—the palest pink is the best endpoint color, not an extremely dark pink. (On your first trial, it is easy to overshoot the endpoint. Although there are some tricks a chemist can try at this point, it is best to do another trial, but this time adding titrant more slowly as you approach the endpoint.)
8. At the endpoint, read the buret again (to two decimal places), and record this as the final volume reading in Table 1. Subtract the two readings; this is the volume of NaOH that is required to neutralize the HCl in the pipetted sample.
9. Refill the buret if necessary, and repeat the titration (Steps 5-8) one or two more times so that two results are within 0.20 mL of each other.

TABLE 1: STANDARDIZATION OF NaOH SOLUTION

Volume of HCl used: <u>20.0</u> mL	Trial 1	Trial 2	Trial 3
Molarity of the HCl: <u>0.955%</u> M			(if needed)
Final buret reading	<u>18.52</u>	<u>18.92</u>	<u>19.51</u>
Initial buret reading	<u>0.01</u>	<u>0.03</u>	<u>0.02</u>
Volume of NaOH used for titration:	<u>18.51</u>	<u>18.89</u>	<u>19.49</u>
Molarity of the NaOH solution: *	<u>0.0469</u>	<u>0.0477</u>	<u>0.046</u>

* From step 11.

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