

L18: CHEMICAL COMPOUNDS OF GREATER COMPLEXITY

IONS

Monoatomic ion: consist of a single atom

Polyatomic ion: made up of a few atoms joined by covalent bonds

POLYATOMIC IONS

- Polyatomic anions are much more common than polyatomic cations.
- Know 2 polyatomic cations
 - ammonium, NH_4^+
 - mercury, Hg_2^{2+}

POLYATOMIC ANIONS (MOST ARE OXOANIONS)

1) Consist of a "core" element

- can be metal or nonmetal (*metals are more common*)
- **+OS**
- Chromium, Cr, and Manganese, Mn are the *only metals* that make polyatomic anions
 - chromate, CrO_4^{2-}
 - dichromate, $\text{Cr}_2\text{O}_7^{2-}$
 - permanganate, MnO_4^-

2) Other atoms are bonded to the "core element"

- can be bonded through single or double covalent bonds
- Usually **-OS**

3) *Most* of the common polyatomic ions are *oxoanions*

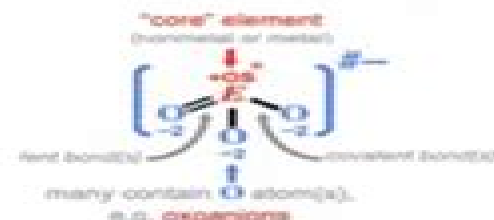
- Oxoanions: contain O (as the "*other atoms*" which are bonded to the "core element" through either single or double bonds
- some oxoanions may also contain various numbers of H atoms

example below

- most of the common polyatomic anions are oxoanions, *except for*:
 - hydroxide ion, OH^-
 - cyanide ion, CN^-

4) the total/net charge of the polyatomic anion will be the difference between:

- the +OS of the "core element"
- the combined -OS of the "other elements"



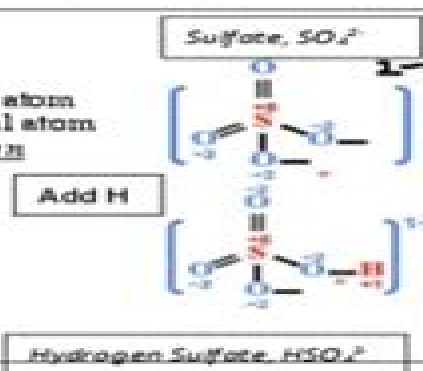
Example: How oxoanions can contain H atoms in their structure

Some oxoanions can also contain H atoms

- In the case of *sulfate*, SO_4^{2-}
- There is a charge of 2- which is a result of 2 unbonded electrons (one on each O atom that is attached by a single bond). This allows each O atom to bond to an additional atom
- When an H atom (OS +1) bonds to one of the oxygens, the ion becomes *hydrogen sulfate*, HSO_4^- , with a charge of 1- because
 - the difference between +OS and -OS becomes smaller
 - there is now only 1 unbonded electron

If we introduce 1 more H atom, the ion becomes a neutral molecule,

Sulfuric acid, H_2SO_4



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