

**2020-2021 Chemical Engineering**  
**Examination 1, August 20, 2020**

Name: \_\_\_\_\_

Name: \_\_\_\_\_

Thermodynamic properties of water vapor at 10. MPa (sat.) are given in the table below in SI units, where  $h$  is in  $\text{kJ kg}^{-1}$ ,  $u$  is in  $\text{kJ kg}^{-1}$ ,  $s$  is in  $\text{kJ kg}^{-1} \text{K}^{-1}$ , and  $c_p$  is in  $\frac{\text{kJ}}{\text{kg K}}$ .

The enthalpy of formation of  $\text{H}_2\text{O}$  is  $-285.83 \frac{\text{kJ}}{\text{mol}}$ .

Multiple-choice questions: 3 questions. Mark the answer considered to be producing the most precise calculation that fits within question. If an answer is not a precise match, it is the closest value to the correct value in the given set of choices.

For questions 1 and 2 use the following information. For the requirements the unit of thermodynamic property listed is as experimentally determined value:  $\frac{\text{kJ}}{\text{kg}}$  or  $\frac{\text{kJ}}{\text{kg K}}$ ,  $\frac{\text{kJ}}{\text{mol}}$  or  $\frac{\text{kJ}}{\text{mol K}}$ .

Water:  $c_p = 1.88 \frac{\text{kJ}}{\text{kg K}}$ ,  $u_{\text{sat}} = 2655.3 \frac{\text{kJ}}{\text{kg}}$ ,  $h_{\text{sat}} = 2776.4 \frac{\text{kJ}}{\text{kg}}$ ,  $s_{\text{sat}} = 6.5958 \frac{\text{kJ}}{\text{kg K}}$ .

1. The specific enthalpy of water vapor at 10 MPa (sat.) is closest to which of the following values?

- |                                       |                                       |                                       |                                       |
|---------------------------------------|---------------------------------------|---------------------------------------|---------------------------------------|
| a. $2655 \frac{\text{kJ}}{\text{kg}}$ | b. $2776 \frac{\text{kJ}}{\text{kg}}$ | c. $2858 \frac{\text{kJ}}{\text{kg}}$ | d. $2981 \frac{\text{kJ}}{\text{kg}}$ |
| e. $3102 \frac{\text{kJ}}{\text{kg}}$ | f. $3225 \frac{\text{kJ}}{\text{kg}}$ | g. $3348 \frac{\text{kJ}}{\text{kg}}$ | h. $3471 \frac{\text{kJ}}{\text{kg}}$ |

2. The unit of enthalpy of formation for the water vapor listed is closest to:

- |  |  |  |  |
|--|--|--|--|
| a. $-285 \frac{\text{kJ}}{\text{mol}}$ | b. $-286 \frac{\text{kJ}}{\text{mol}}$ | c. $-287 \frac{\text{kJ}}{\text{mol}}$ | d. $-288 \frac{\text{kJ}}{\text{mol}}$ |
| e. $-289 \frac{\text{kJ}}{\text{mol}}$ | f. $-290 \frac{\text{kJ}}{\text{mol}}$ | g. $-291 \frac{\text{kJ}}{\text{mol}}$ | h. $-292 \frac{\text{kJ}}{\text{mol}}$ |

3. Estimate the following value using the information:  $h_{\text{sat}} = 2776.4 \frac{\text{kJ}}{\text{kg}}$ ,  $u_{\text{sat}} = 2655.3 \frac{\text{kJ}}{\text{kg}}$ .

of water	of water	of water
at 10 MPa	at 10 MPa	at 10 MPa
liquid	liquid	liquid
solid	solid	solid

4. Estimate the value of enthalpy of water vapor:

- |                                       |                                       |                                       |
|---------------------------------------|---------------------------------------|---------------------------------------|
| a. $2655 \frac{\text{kJ}}{\text{kg}}$ | b. $2776 \frac{\text{kJ}}{\text{kg}}$ | c. $2858 \frac{\text{kJ}}{\text{kg}}$ |
| d. $2981 \frac{\text{kJ}}{\text{kg}}$ | e. $3102 \frac{\text{kJ}}{\text{kg}}$ | f. $3225 \frac{\text{kJ}}{\text{kg}}$ |

5. Estimate the specific heat capacity of water vapor. It is a diatomic gas at 10 MPa (sat.) and is assumed to have a constant specific heat capacity. What is the most precise estimate of the specific heat capacity of water vapor?

- |   |   |   |   |
|---|---|---|---|
| a. $1.88 \frac{\text{kJ}}{\text{kg K}}$ | b. $1.89 \frac{\text{kJ}}{\text{kg K}}$ | c. $1.90 \frac{\text{kJ}}{\text{kg K}}$ | d. $1.91 \frac{\text{kJ}}{\text{kg K}}$ |
|---|---|---|---|

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