

Determination of the Molar Mass of an Unknown Acid

Introduction

In this experiment you will determine the molar mass of an unknown acid by titration with the sodium hydroxide solution you prepared and standardized in the previous experiment. To do so, you will need to prepare a solution of known volume, using a known mass of the solid unknown acid as the solute. By titrating the unknown acid solution against a standardized base, you will be able to calculate the molarity of the acid solution to a high degree of precision. Multiplying the molarity by the volume of solution will yield the number of moles of acid in solution; dividing this into the mass of acid used in the solution preparation will indicate the molar mass.

Your unknown acid is a polyprotic acid, meaning that it contains more than one acidic (reactive) hydrogen per molecule. A diprotic acid, which is an acid containing 2 acidic hydrogens per molecule, can be represented as H_2A . Similarly, a triprotic acid (one that contains 3 acidic hydrogens per molecule) can be represented as H_3A .

The general steps you will complete in this experiment (solution preparation followed by titration) are very similar to the previous experiment. Only minimal detail will be given in the procedure. You will be expected to use the knowledge you obtained from last week's experiment to complete this week's experiment successfully. If you are uncertain of any steps that are required, please refer to the [Standardization of a NaOH Solution](#) experiment.

Pre-Lab Assignment

Summary: Give a brief overview of the experiment in your lab notebook using blue or black pen. Your summary should be 2-4 sentences long and written in your own words. Be sure to include any chemicals used in the experiment.

Data Tables: Construct your data tables in your lab notebook using **blue or black pen** (no pencil) and a **straightedge** (ruler). The format of the data tables will not be provided, since they are similar to those of the previous experiment. You should create your data table based on the data you expect to collect in the experiment. Your data tables must be properly constructed in your lab notebook **prior** to lab. Be certain to leave a space to record your unknown code.

Pre-lab question: Answer the following question in your lab notebook. Answers to prelab questions should not be written on the same page as the data tables. Your response should be submitted prior to the beginning of the lab.

A 0.2535 M solution of an unknown ionic compound is prepared by adding 5.255 g of the compound to a 125 mL volumetric flask and adding deionized water to the line. What is the molar mass of the ionic compound? Show your work and include the correct number of significant figures.

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