

Name: Lily Quising

1. Calculate the volume of the standard NaOH solution needed to do the titration.

$$M_1 V_1 = M_2 V_2 \rightarrow 3.0 M NaOH \times V_1 = 500.00 mL \times 0.18 M$$

$$V_1 = \frac{500.00 mL \times 0.18 M}{3.0 M NaOH}$$

$$V_1 = 30 mL$$

2. Concentration of
- H_2SO_4
- solution:
- 0.1526 M

3. Standardization of NaOH solution:

	Titration 1	Titration 2	Titration 3	Titration 4
Volume of H_2SO_4	20.00 mL	20.00 mL	20.00 mL	20.00 mL
Initial buret reading	0.00 mL	0.50 mL	0.75 mL	1.50 mL
Final buret reading	32.00 mL	33.29 mL	33.57 mL	34.90 mL
Volume of base used	32.00 mL	32.79 mL	32.82 mL	33.40 mL
Depth of pink color	Deep pink	Light pink	Light pink	Light pink

4. Calculate the molarity of the NaOH solution. Show the calculation for at least one titration.



$$M = ? \quad M = 0.1526 M$$

$$V = 33.40 mL \quad V = 20.00 mL$$

$$\text{moles of } H_2SO_4 = 0.1526 M \times 0.0200 L$$

$$= 0.003052 \text{ mol } H_2SO_4$$

$$0.003052 \text{ mol } H_2SO_4 \times \frac{2 \text{ mol NaOH}}{1 \text{ mol } H_2SO_4} = 0.006104 \text{ mol NaOH}$$

$$\text{Molarity of NaOH} = \frac{0.006104 \text{ mol NaOH}}{0.03340 L NaOH} = 0.1828 M (NaOH)$$

	Titration 1	Titration 2	Titration 3	Titration 4
Molarity of NaOH	0.1808 M	0.1862 M	0.1860 M	0.1828 M

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